DBMS Quiz-1

mohan@jnnce.ac.in Switch account

* Indicates required question

DBMS Quiz-1

In ER diagram, a rectangle is used to represent _____. * 1 point

- Entity Type
- Attribute
- Relationship
- Cardinality

Entity types that do not have key attributes of their own are called _____. * 1 point

- weak entity types
- strong entity types
- no key relation types

Educational Degrees of a Person is an example of _____. * 1 point

- Composite Attribute
- Multivalued Attribute
- Derived Attribute
- Simple Attribute

Which of the following is NOT the description of the Database : * 1 point

- MetaData
- Extension
- Catalog
- Schema

A state that satisfies the structure and constraints of the Database is a: * 1 point

- Entity State
- Schema State
- Valid State
- Initial State

The _____ specifies whether the existence of an entity depends on its being related to another entity via the relationship type. * 1 point

- participation constraint
- domain constraint
- key constraint
- semantic constraint

The number of Attributes in a Relation is called as : * 1 point

- Degree of a Relation
- Constraints of a Relation
- Meta Data of a Relation
- Cardinality of a Relation

The capacity to change the conceptual schema without having to change the external schemas and their application programs. * 1 point

- Logical Data Independence
- Physical Data Independence

The Architecture of the Database can be viewed as: * 1 point

- Four Levels
- Two Levels
- Three Levels
- One Level

Values for an attribute of a Tuple is selected from a _____. * 1 point

- Foreign Key
- Domain
- Another Relation
- Database

In DBMS terminology a Column is : * 1 point

- Degree
- Relation
- Tuple
- Attribute

The number of Rows in a Table is called as : * 1 point

- Degree of a Relation
- Constraints of a Relation
- Cardinality of a Relation
- Meta Data of a Relation

User responsible for authorizing access to the database is _____. * 1 point

- End User
- Database Administrator
- Database Designers
- System Analysts

Address Attribute is an example for _____. * 1 point

- Composite Attribute
- Multivalued Attribute

- ◯ Derived Attribute
- ◯ Simple Attribute

The relationship type that relates a weak entity type to its owner is called as _____ of the weak entity type. * 1 point

- ◯ identifying relationship
- ◯ underlying relationship
- ◯ weak relationship
- ◯ binary relationship

The capacity to change the internal schema without having to change the conceptual schema. * 1 point

- ◯ Logical Data Independence
- ◯ Physical Data Independence

No Two Student Entities can have same USN value. This Statement in DBMS is a: * 1 point

- ◯ Constraint
- ◯ Declaration
- ◯ Meta Data
- ◯ Domain

In DBMS terminology a Row is a : * 1 point

- ◯ Degree
- ◯ Relation
- ◯ Tuple
- ◯ Attribute

The capacity to change the schema at one level of a database system without having to change the system at the next higher level is: * 1 point

- ◯ Meta Data Independence
- ◯ Data Independence
- ◯ Program Independence
- ◯ Schema Independence

The Database Schema changes very rarely. * 1 point

- ◯ True
- ◯ False

The actual data stored in a Database at a particular moment of time is called as _____. * 1 point

- ◯ Extension
- ◯ Meta Data
- ◯ Intension
- ◯ Database State

Number of Employees in a Company is an example for _____. * 1 point

- ◯ Composite Attribute
- ◯ Multivalued Attribute
- ◯ Derived Attribute
- ◯ Simple Attribute

The Database State does not change every time the database is updated. * 1 point

- ◯ True
- ◯ False

The cardinality ratio and participation constraint together called as _____. * 1 point

- ◯ key constraint
- ◯ structural constraint
- ◯ domain constraint
- ◯ entity constraint

Find the Odd one: * 1 point

- ◯ Intension
- ◯ Table
- ◯ Relation
- ◯ Entity Type
- ◯ Attribute

The view of total Database content is: * 1 point

- ◯ External View
- ◯ Internal View
- ◯ Physical View
- ◯ Conceptual View

DBMS will help in Controlling Redundancy * 1 point

- ◯ True
- ◯ False

The cardinality ratio for binary relationship specifies the _____ number of relationship instances that an entity can participate in. * 1 point

- ◯ minimum
- ◯ maximum

DBMS will provides persistent storage for program Objects * 1 point

- ◯ True
- ◯ False

In a Application both Date of Birth and Age of a person is used. Which of the Following is Correct: * 1 point

- ◯ Both Date of Birth and Age are Derived Attributes
- ◯ Both Date of Birth and Age are Stored Attributes

Both Date of Birth and Age are Stored Attributes

Date of Birth is Derived Attribute and Age is Stored Attribute

Date of Birth is Stored Attribute and Age is Derived Attribute

A copy of your responses will be emailed to the address you provided.

Back Submit Page 2 of 2 Clear form

Never submit passwords through Google Forms.

DBMS (2021-22) Quiz-1

Send

Questions | Responses 67 | Settings

Total points: 30

67 responses

View in Sheets

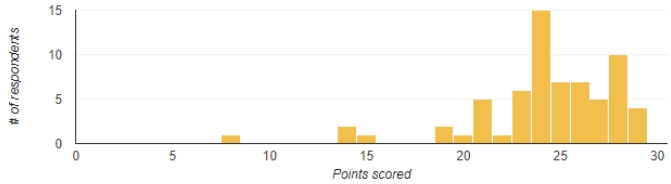
Accepting responses

Summary | Question | Individual

Insights

Average	Median	Range
24.21 / 30 points	24 / 30 points	8 - 29 points

Total points distribution



Frequently missed questions

Question	Correct responses
Find the Odd one:	14 / 67
Which of the following is NOT the description of the Database :	32 / 67
Number of Employees in a Company is an example for _____.	14 / 67

Scores

Release scores

Email	Score / 30	Score released
leelagowda17028@gmail.com	24	Nov 18 3:16 PM
jsuchitrarbhatagni@gmail.com	14	Nov 18 3:16 PM
shobhitarajanna3@gmail.com	8	Nov 18 3:16 PM
4jn20cs401chandan@gmail.com	26	Nov 18 3:16 PM
mohan@jnnce.ac.in	25	Nov 18 3:16 PM
4jn19cs096sinchananoolee@gmail.com	21	Nov 18 3:16 PM
shoaibsayed0@gmail.com	14	Nov 18 3:16 PM
4jn19cs066pavanakp@gmail.com	24	Nov 18 3:16 PM

USN

67 responses

4JN20CS404
4JN19CS106
4JN18CS090
4JN20CS401
4JN
4JN19CS096

4JN19CS090

4JN19CS066

4JN19CS100

Name

67 responses

Leelavathi s

Suchitra R Bhat Agni

Shobhita G R

Chandan Singh

Mohan

Sinchana Noolee

Shoaib kaleem sayad

Pavana K P

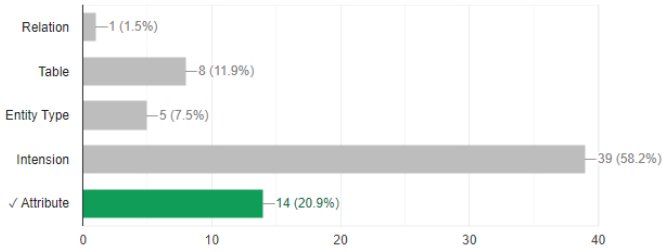
Sneha HM

DBMS Quiz-1

Find the Odd one:

Copy

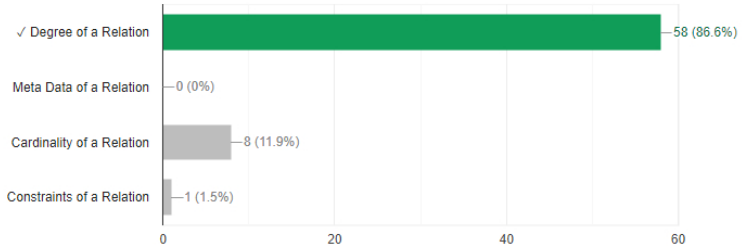
14 / 67 correct responses



The number of Attributes in a Relation is called as :

Copy

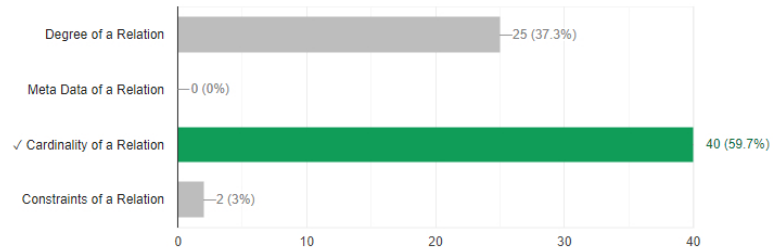
58 / 67 correct responses



The number of Rows in a Table is called as :

Copy

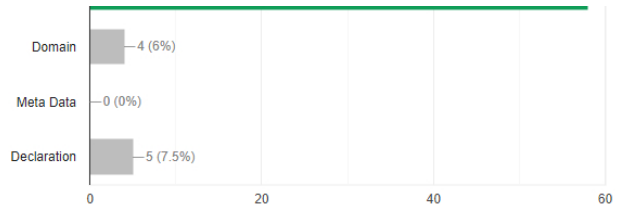
40 / 67 correct responses



No Two Student Entities can have same USN value. This Statement in DBMS is a:

Copy

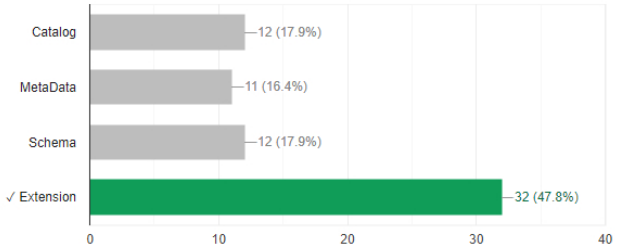
58 / 67 correct responses



Which of the following is NOT the description of the Database :

Copy

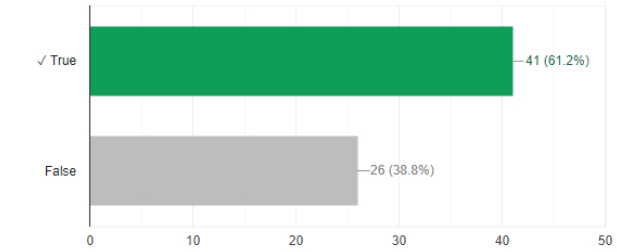
32 / 67 correct responses



The Database Schema changes very rarely.

Copy

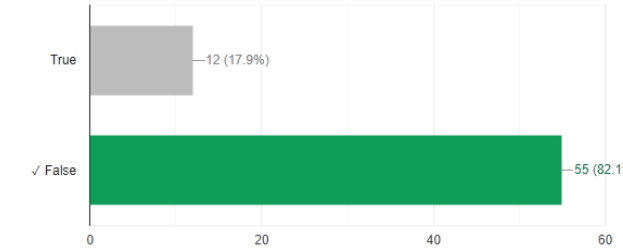
41 / 67 correct responses



The Database State does not change every time the database is updated.

Copy

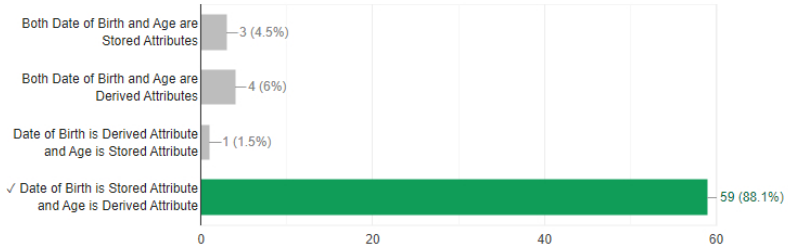
55 / 67 correct responses



In a Application both Date of Birth and Age of a person is used. Which of the Following is Correct:

Copy

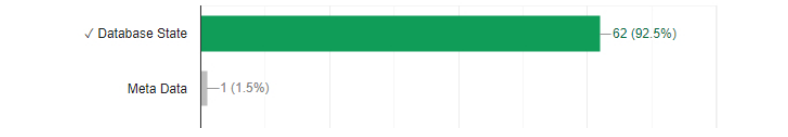
59 / 67 correct responses



The actual data stored in a Database at a particular moment of time is called as _____.

Copy

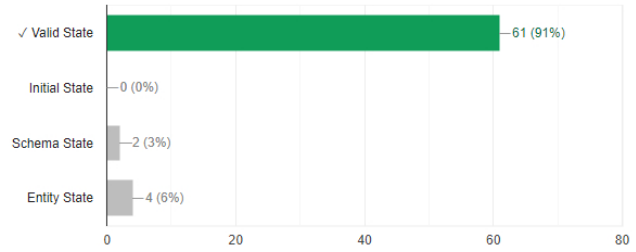
62 / 67 correct responses



A state that satisfies the structure and constraints of the Database is a:

Copy

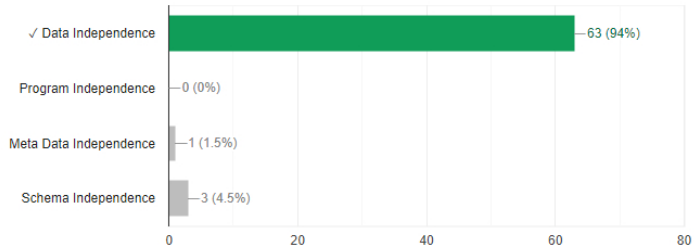
61 / 67 correct responses



The capacity to change the schema at one level of a database system without having to change the system at the next higher level is:

Copy

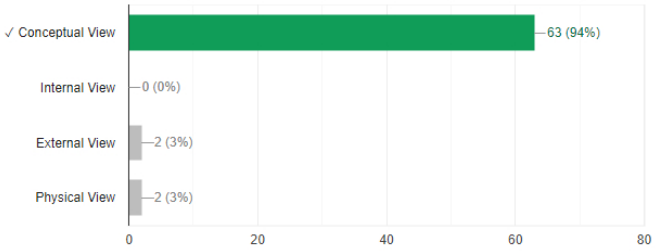
63 / 67 correct responses



The view of total Database content is:

Copy

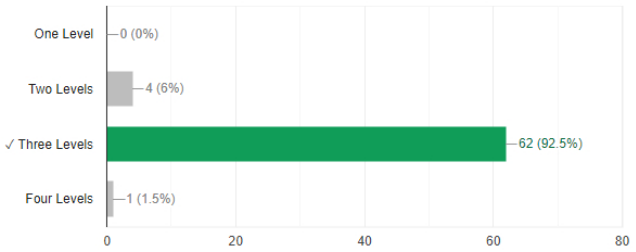
63 / 67 correct responses



The Architecture of the Database can be viewed as:

Copy

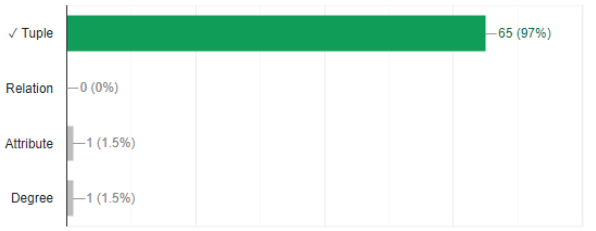
62 / 67 correct responses



In DBMS terminology a Row is a :

Copy

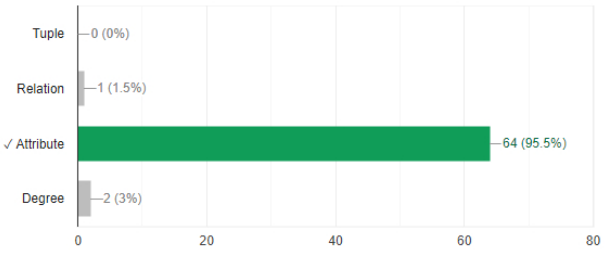
65 / 67 correct responses



In DBMS terminology a Column is :

Copy

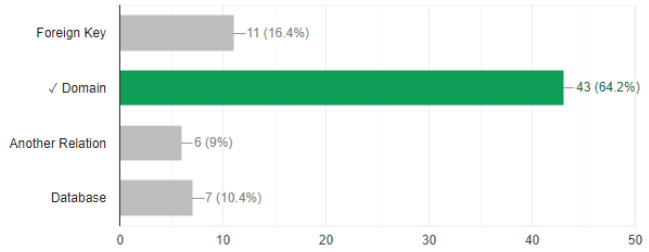
64 / 67 correct responses



Values for an attribute of a Tuple is selected from a _____.

Copy

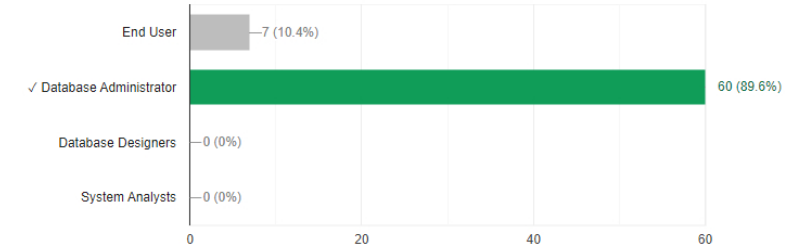
43 / 67 correct responses



User responsible for authorizing access to the database is _____.

Copy

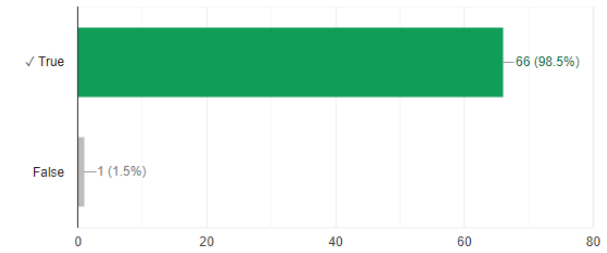
60 / 67 correct responses



DBMS will help in Controlling Redundancy

Copy

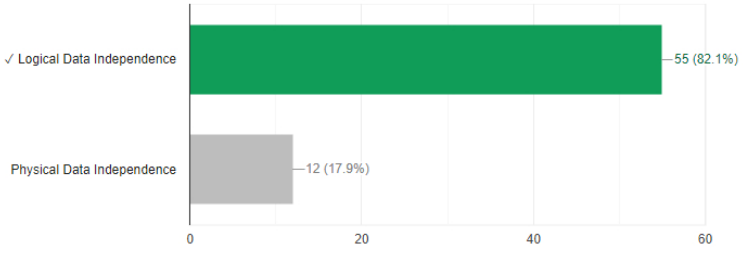
66 / 67 correct responses



The capacity to change the conceptual schema without having to change the external schemas and their application programs.

Copy

55 / 67 correct responses

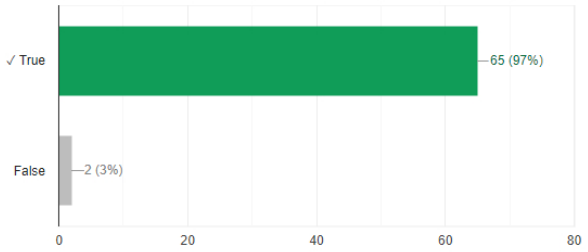


DBMS will provides persistent storage for program Objects

Copy

DBMS will provides persistent storage for program objects

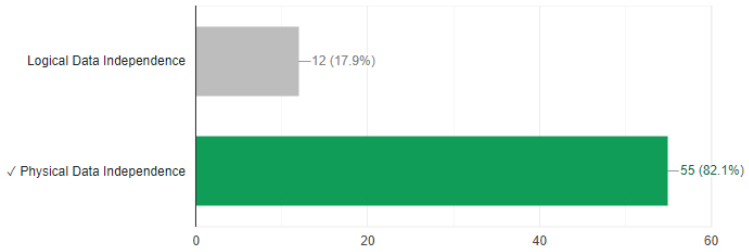
65 / 67 correct responses



The capacity to change the internal schema without having to change the conceptual schema.

Copy

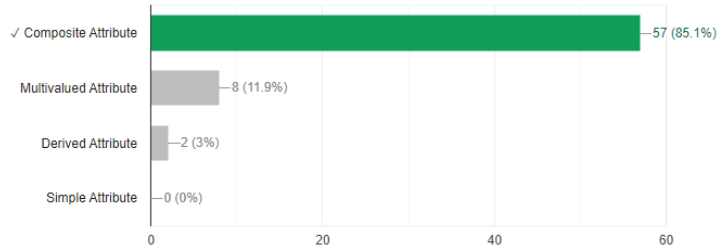
55 / 67 correct responses



Address Attribute is an example for _____.

Copy

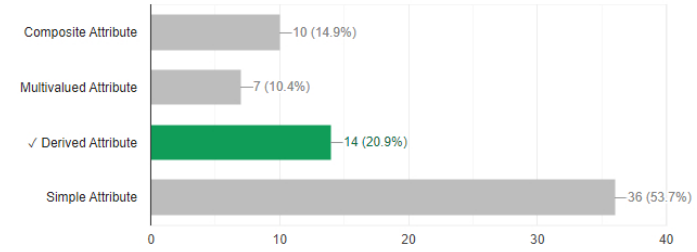
57 / 67 correct responses



Number of Employees in a Company is an example for _____.

Copy

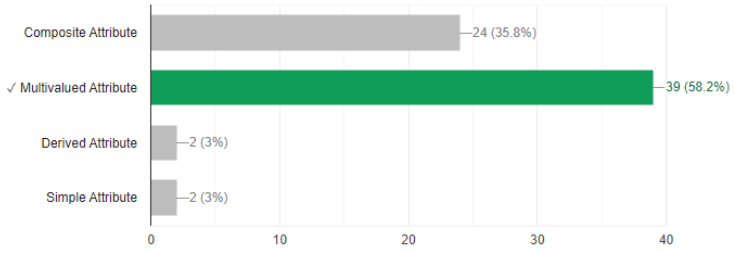
14 / 67 correct responses



Educational Degrees of a Person is an example of _____.

Copy

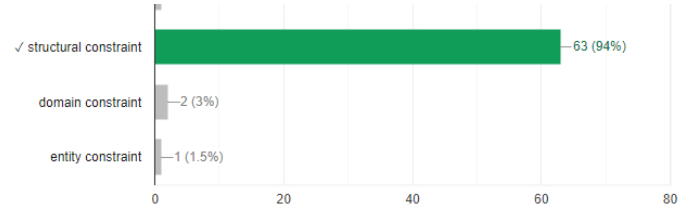
39 / 67 correct responses



The cardinality ratio and participation constraint together called as _____.

Copy

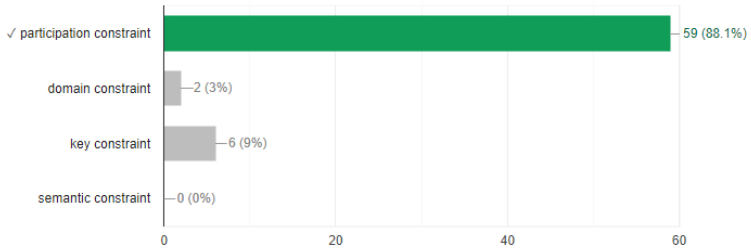
63 / 67 correct responses



The _____ specifies whether the existence of an entity depends on its being related to another entity via the relationship type.

Copy

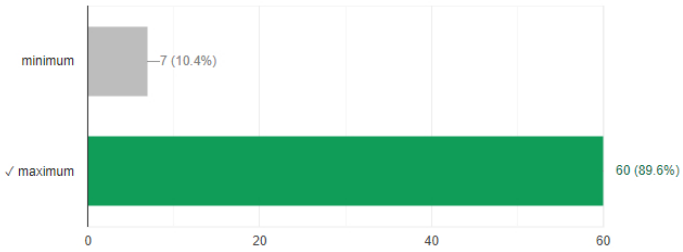
59 / 67 correct responses



The cardinality ratio for binary relationship specifies the _____ number of relationship instances that an entity can participate in.

Copy

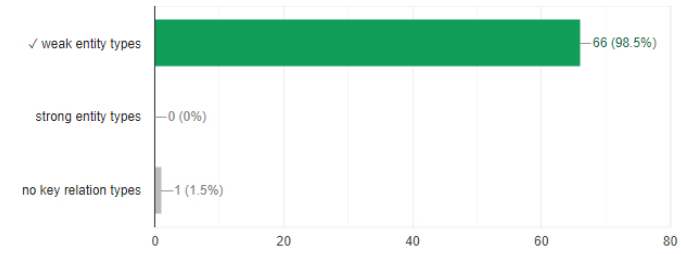
60 / 67 correct responses



Entity types that do not have key attributes of their own are called _____.

Copy

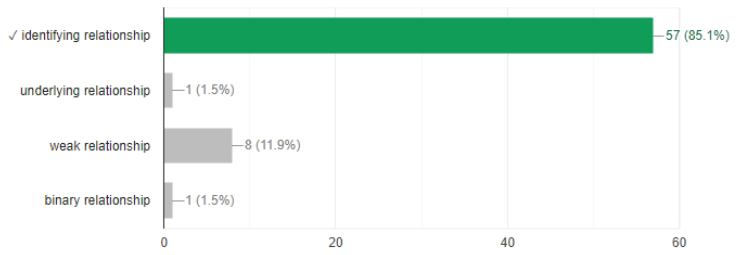
66 / 67 correct responses



The relationship type that relates a weak entity type to its owner is called as _____ of the weak entity type.

Copy

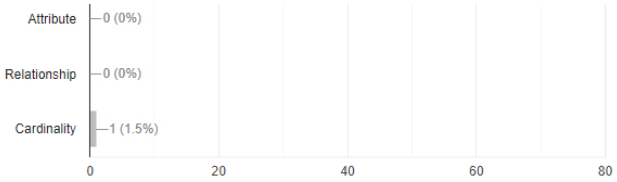
57 / 67 correct responses



In ER diagram, a rectangle is used to represent _____.

Copy

66 / 67 correct responses


Attribute
0 (0%)
Relationship
0 (0%)
Cardinality
1 (1.5%)
0
20
40
60
80

(2) WhatsApp | Inbox (5,585) - vedanandade@jn | DAA_Mod4 Quiz - Google Forms | DAA_Mod4_Quizz | DAA_Mod4 Quiz - Google Forms

docs.google.com/forms/d/1G1qTJ5McwKJS_JSUlBtwLFMT8mVxrQTPNMxKvXl16w/edit#responses

https://www.gmail.c... | jnnce.ac.in/cse/resu... | Gmail | JNNCE CSE | 01204325 - Data C... | Gmail | www.google.com | untitled | Attn. SSD owners -... | Google Meet | Certificate Maker –...

DAA_Mod4 Quiz

Send

Questions | Responses 70 | Settings | Total points: 10

70 responses

View in Sheets

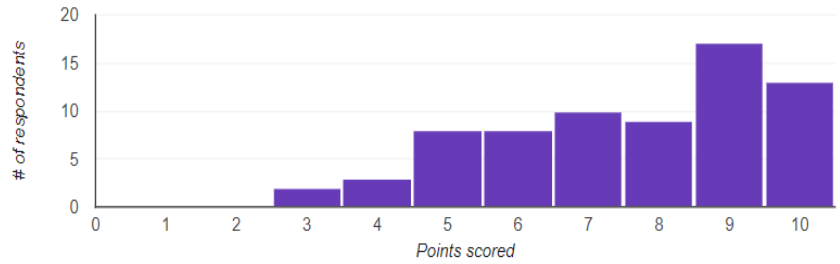
Accepting responses

Summary | Question | Individual

Insights

Average	Median	Range
7.59 / 10 points	8 / 10 points	3 - 10 points

Total points distribution



30°C Mostly sunny

Search

ENG US

6:54 01-09-2023

Subject:ETM (Mrs. Pushpa R N)

#	Question	Question Type	Question Accuracy	Average Time per Question (mm:ss)	Correct	Yet to be graded	Partially correct	Incorrect	Ungraded	Unattempted
1	Which of the following is true about Princip	Multiple Choice	86%	00:19	32	0	0	3	0	2
2	Which of the following best describes the P	Multiple Choice	86%	00:19	32	0	0	3	0	2
3	The principles of management serve as a ge	Multiple Choice	84%	00:15	31	0	0	4	0	2
4	components of Direction in management	Multiple Choice	95%	00:10	35	0	0	0	0	2
5	Management skills involves ?	Multiple Choice	95%	00:07	35	0	0	0	0	2
6	Management satisfies what characteristics	Multiple Choice	76%	00:15	28	0	0	7	0	2
7	Management is	Multiple Choice	92%	00:10	34	0	0	2	0	1
8	The first and foremost function of managem	Multiple Choice	89%	00:11	33	0	0	2	0	2
9	Goals, aims, purposes, missions and target	Multiple Choice	84%	00:14	31	0	0	4	0	2
10	The process of establishing a time sequenc	Multiple Choice	78%	00:15	29	0	0	7	0	1
			86%	02:02	320	0	0	32	0	18

Quizizz

View Player Data | View Time Data | View Summary

#	Question	Question Type	Question Accuracy	Average Time per Question (mm:ss)
1	Which of the following is true about Principle	Multiple Choice	86%	00:19
2	Which of the following best describes the Prir	Multiple Choice	86%	00:19
3	The principles of management serve as a gene	Multiple Choice	84%	00:15
4	components of Direction in management	Multiple Choice	95%	00:10
5	Management skills involves ?	Multiple Choice	95%	00:07
6	Management satisfies what characteristics of	Multiple Choice	76%	00:15
7	Management is	Multiple Choice	92%	00:10
8	The first and foremost function of managemer	Multiple Choice	89%	00:11
9	Goals, aims, purposes, missions and target is a	Multiple Choice	84%	00:14
10	The process of establishing a time sequence f	Multiple Choice	78%	00:15
			86%	2:2

Quizizz

Rank	First Name	Last Name	Attempt #	Accuracy	Score	Correct
1	Chandan	S 4JN21CS035	10	100%	10420	10
2	Jeevitha.P	4JN21CS062	10	100%	10240	10
3	Dhanyashree	B	10	100%	9970	10
4	Abhishek	H J.4JN21CS003	10	100%	9810	10
5	4JN21CS051	GANAVI A C	10	100%	9420	10
6	Inchara	D, 4JN21CS058	10	100%	9390	10
7	Anukeerthana	MB,4JN21CS02	10	100%	9280	10
8	4JN21CS030	Bharath C	10	100%	9160	10
9	4JN21CS026		10	100%	8920	10
10	Aishwarya	S G	10	100%	8860	10
11	Gurukiran	K A 4JN21CS054	10	100%	8860	10
12	4JN21CS004		10	100%	8830	10
13	4JN21CS014		10	100%	8620	10
14	Adamya		10	100%	8560	10
15	ABHISHEK	K S 4JN21CS006*	10	90%	8170	9
16	Hongirana	M 4JN21CS057	10	100%	8000	10
17	Aishwarya	K P	10	90%	7840	9
18	4JN21CS029	Atiya	10	100%	7760	10
19	Disha	D 4JN21CS046	10	90%	7710	9
20	Iqra	Fayaz 4JN21CS060	10	90%	7620	9
21	Arpitha	V 4JN21CS027	10	90%	7530	9
22	4JN21CS020		10	90%	7480	9
23	4jn21cs008		10	90%	7440	9
24	Ananya	S CSE 4JN21CS022	10	90%	7340	9
25	Bhumika.H.C	4JN21CS032	10	90%	7140	9
26	Hiranmayi	S 4JN21CS056*	10	80%	7000	8
27	Abhishek	S	10	90%	6920	9
28	4JN21CS018	Akash L Naik	10	80%	6750	8
29	Aishwarya	4JN21CS013	10	80%	6600	8
30	Harshitha	M 4JN21CS055	10	80%	6560	8
31	ABHISHEK	K S 4JN21CS006	10	70%	6510	7
32	Dimple		10	80%	6460	8
33	Aishwarya	Patil HM	10	80%	6440	8
34	4JN21CS048,Divya		10	80%	6390	8

Sample quiz question paper answered by the student

DSP(18EC52) 2021-22 (odd)

2021-22 (odd)

Name: Pratima Meghashyam Naik

USN: 4JN19ETO24

J.N.N COLLEGE OF ENGINEERING, SHIVAMOGGA
DEPARTMENT OF ELECTRONICS AND TELECOMMUNICATION ENGINEERING

QUIZ

Semester: 5-CBCS
Subject: DSP (18EC52)

Date: 14 Jan 2022
Faculty: Mr.Harisha S B

1 Determine the number of complex additions required for 32 direct computations of DFT.
a) 240 b) 56 c) 992 d) 854

2 Determine the number of complex multiplications for the 8-point Radix-2 FFT.
a) 32 b) 12 c) 80 d) 4

3 The Z-transform of the function y(n) = x(n) + y(n - 1) is:
a) z/ z + 1 b) z/ 2z c) z/ z - 1 d) z - 1/z

4 The window technique whose main lobe width is 12pi/N is called:
a) Hamming window b) Blackmann window c) Kaiser window d) Rectangular window.

5 If x(n) is a real sequence and X(k) is its N-point DFT, then which of the following is true?
a) X(N-k)=X(-k) b) X(N-k)=X*(k) c) X(-k)=X*(k) d) All of the mentioned

6 In Overlap save method of long sequence filtering, what is the length of the input sequence block?
a) L+M+1 b) L+M c) L+M-1 d) None of the mentioned

7 For a decimation-in-time FFT algorithm, which of the following is true?
a) Both input and output are in order b) Both input and output are shuffled c) Input is shuffled and output is in order d) Input is in order and output is shuffled

8 What is the output of the single stage lattice filter if x(n) is the input?
a) x(n)+Kx(n+1) b) x(n)+Kx(n-1) c) x(n)+Kx(n-1)+Kx(n+1) d) Kx(n-1)

9 If M and N are the orders of numerator and denominator of rational system function respectively, then how many memory locations are required in direct form-II realization of that IIR filter?
a) M+N+1 b) M+N c) Min [M,N] d) Max [M,N]

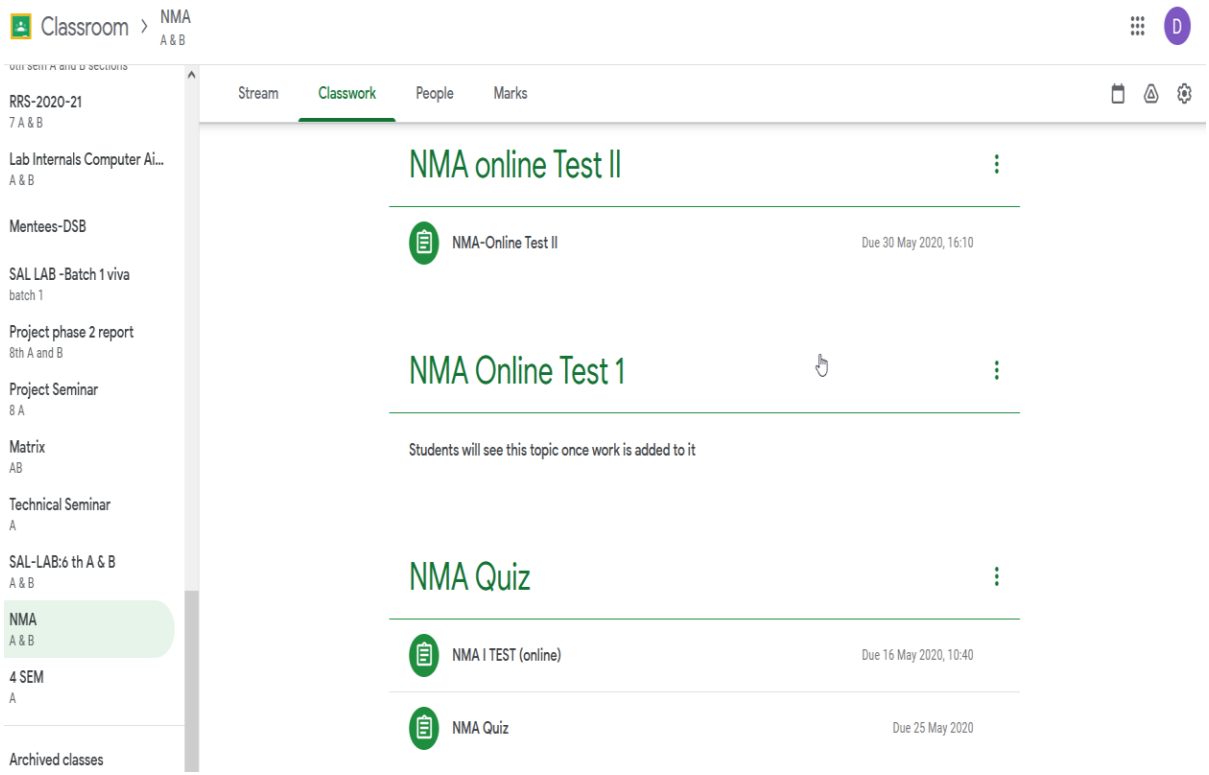
10 In IIR Filter design by the Bilinear Transformation, the Bilinear Transformation is a mapping from
a) Z-plane to S-plane b) S-plane to Z-plane c) S-plane to J-plane d) J-plane to Z-plane

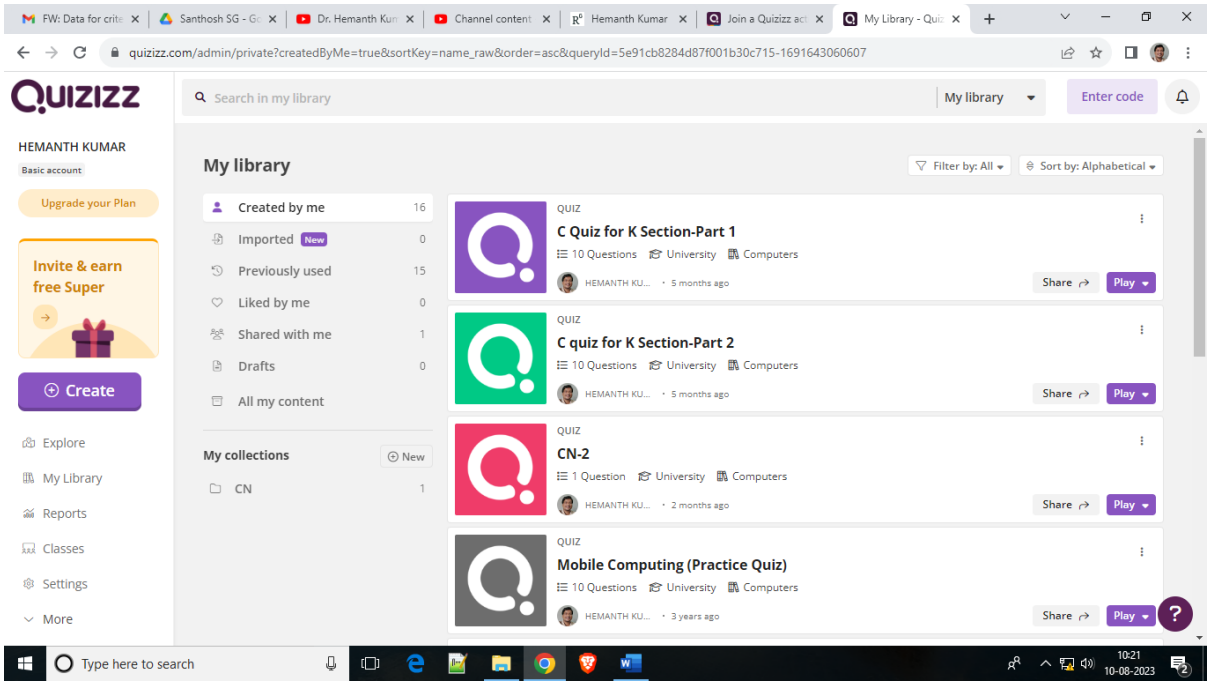
$x(n)$

$X(32) = \sum^{31} x(n) w_n^1$

$1 + z^{-1}$ $\quad z - 1^{-z}$ $\quad \frac{z}{z} - 1$

Subject: Numerical Methods and its application(17CV663)


Classroom
NMA
A & B
RRS-2020-21
7 A & B
Lab Internals Computer Ai...
A & B
Mentees-DSB
SAL LAB -Batch 1 viva
batch 1
Project phase 2 report
8th A and B
Project Seminar
8 A
Matrix
AB
Technical Seminar
A
SAL-LAB:6 th A & B
A & B
NMA
A & B
4 SEM
A
Archived classes
Stream
Classwork
People
Marks
NMA online Test II
NMA-Online Test II
Due 30 May 2020, 16:10
NMA Online Test 1
Students will see this topic once work is added to it
NMA Quiz
NMA I TEST (online)
Due 16 May 2020, 10:40
NMA Quiz
Due 25 May 2020


quizizz.com/admin/private?createdByMe=true&sortKey=name_raw&order=asc&queryId=5e91cb8284d87f001b30c715-1691643060607
QUIZIZZ
Search in my library
My library
Enter code
HEMANTH KUMAR
Basic account
Upgrade your Plan
Invite & earn free Super
Create
Explore
My Library
Reports
Classes
Settings
More
My library
Filter by: All
Sort by: Alphabetical
Created by me 16
Imported New 0
Previously used 15
Liked by me 0
Shared with me 1
Drafts 0
All my content
My collections
New
CN 1
QUIZ
C Quiz for K Section-Part 1
10 Questions University Computers
HEMANTH KU... 5 months ago
Share
Play
QUIZ
C quiz for K Section-Part 2
10 Questions University Computers
HEMANTH KU... 5 months ago
Share
Play
QUIZ
CN-2
1 Question University Computers
HEMANTH KU... 2 months ago
Share
Play
QUIZ
Mobile Computing (Practice Quiz)
10 Questions University Computers
HEMANTH KU... 3 years ago
Share
Play
Type here to search
10:21
10-08-2023

Quiz on Artificial Intelligence and Machine Learning (18CS753)

Quiz on Artificial Intelligence and Machine Learning (18CS753)

mohan@jnnce.ac.in Switch account

* Indicates required question

Email *

Your email

USN *

Your answer

Name *

Your answer

Email *

Your answer

Quiz Questions

English sentences grammar checking comes under * 1 point

- Natural Language Processing
- Truth Maintanance Systems
- Induction
- Predicate Logic

The solution of the Travelling Salesman Problem is about finding a _____. * 1 point

- path
- state
- number
- accuracy

Single value attributes means * 1 point

- The attribute which consider unique value
- The attribute which consider multiple values
- The attribute which consider no values
- None of the options

Which is the First Stage in Natural Language Processing * 1 point

- Morphological Analysis
- Syntactic Analysis
- Semantic Analysis
- Discourse Integration
- Pragmatic Analysis

Identify the Expert Task * 1 point

- Speech Recognition
- Chess
- Medical Diagnosis

Who is Father of AI ? * 1 point

- Alan Turing
- John McCarthy
- Dennis Ritche
- Elaine Rich

Rules of the production system are of form : * 1 point

- x -> y
- xy
- x = y
- x <- y

Which one is considered as branch of Artificial Intelligence? * 1 point

- Machine Learning
- Database
- Network Design
- Java Programming

Rote learning is a _____. * 1 point

- memorization technique based on repetition
- evaluation technique based on computation
- storage technique based on data
- parsing technique based on tree

In Playing Chess, Goal Position is _____. * 1 point

- any position in which the opponent does not have a legal move and opponent king is under attack
- to save our king from attack
- any position where our king is not under attack
- To clear all the opponent Pawns

Which one is a Mundane Task? * 1 point

- Speech Recognition
- Chess

- Chess
- Medical Diagnosis

Water-Jug Problem solution is represented using * 1 point

- State Space Tree
- Cyclic Graph
- Blocks
- Predicates

Which Test is used to determine whether a machine can think? * 1 point

- Turing Test
- BFS
- DFS
- Fitness Test

Which one is a Formal Task? * 1 point

- Speech Recognition
- Chess
- Medical Diagnosis

Which Stage of Natural Language Processing refer to preceding sentences? * 1 point

- Morphological Analysis
- Syntactic Analysis
- Semantic Analysis
- Discourse Integration
- Pragmatic Analysis

In _____ linear sequences of words are transformed into structures. * 1 point

- Morphological Analysis
- Syntactic Analysis
- Semantic Analysis
- Discourse Integration
- Pragmatic Analysis

A good control strategy must * 1 point

- cause motion
- lead to a solution
- systematic
- All the options

Artificial Intelligence is about_____. * 1 point

- Putting your intelligence into Computer
- Programming with your own intelligence
- Playing a Game
- Making a Machine intelligent

Intelligence requires knowledge * 1 point

- True
- False

A good system for the representation of knowledge in a particular domain should possess how many properties * 1 point

- 4
- 5
- 3
- 6

A copy of your responses will be emailed to the address you provided.

Submit Page 1 of 1 Clear form

Never submit passwords through Google Forms.

Quiz on AIML

Send

Questions | Responses 81 | Settings

Total points: 20

81 responses

View in Sheets

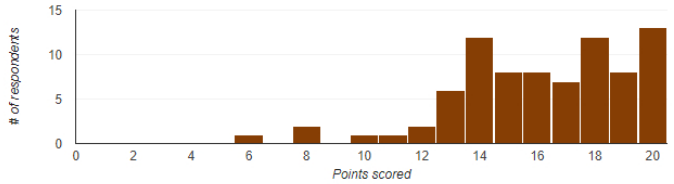
Accepting responses

Summary | Question | Individual

Insights

Average	Median	Range
16.15 / 20 points	16 / 20 points	6 - 20 points

Total points distribution



Frequently missed questions

Question	Correct responses
A good control strategy must	37 / 81
Which Stage of Natural Language Processing refer to preceding sentences?	36 / 81

Scores

Release scores

Email	Score / 20	Score released
spurthim0215@gmail.com	16	Jan 3 10:02 AM
mahankushaacademy@gmail.com	8	Jan 3 10:02 AM
bhumikasgaler42@gmail.com	10	Jan 3 10:02 AM
rharishma921@gmail.com	16	Jan 3 10:02 AM
ashakumarjadav2019@gmail.com	12	Jan 3 10:02 AM
swamysandy@jnnce.ac.in	8	Jan 3 10:02 AM
preethijadavk@gmail.com	18	Jan 3 10:02 AM
sangeethams134@gmail.com	17	Jan 3 10:02 AM

USN

81 responses

4
4JN19EC077
4JN19EC091
4JN18ME027
4JN19ET010
4JN19ET013
4JN19ET025

4JN19EC078

4JN19EC083

Name

81 responses

Sandeep swamy P S

Spurthi M

Mahan

Bhumika s g

Harishma R

Raksha

Preethi Jadav K

Sangeetha M S

Shreenidhi H M

Email

81 responses

spurthim0215@gmail.com

mahankushaacademy@gmail.com

bhumikasgaler42@gmail.com

rharishma921@gmail.com

ashakumarjadav2019@gmail.com

swamysandy@jnnce.ac.in

preethijadavk@gmail.com

sangeethams134@gmail.com

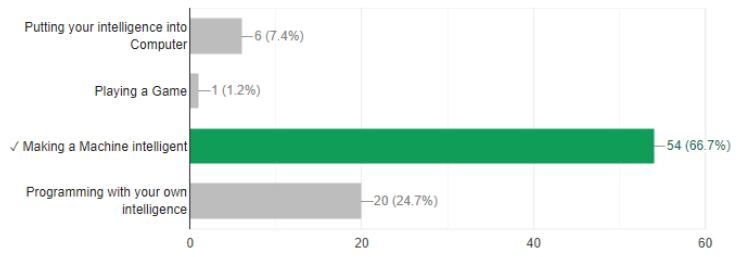
shreenidhihm02@gmail.com

Quiz Questions

Artificial Intelligence is about_____.

Copy

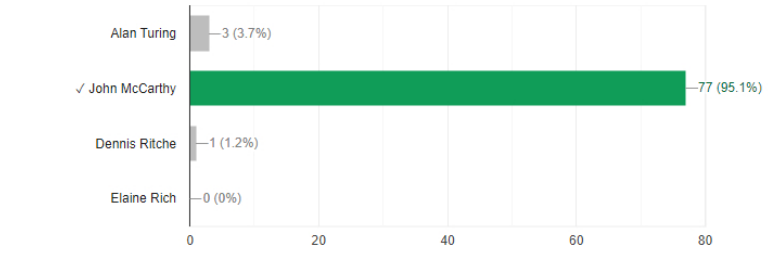
54 / 81 correct responses



Who is Father of AI ?

Copy

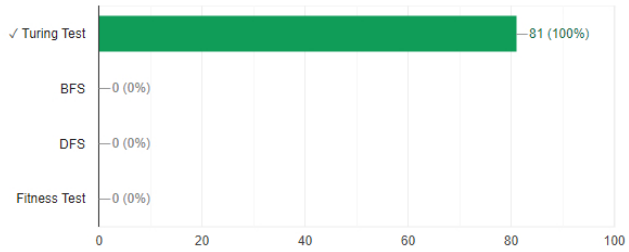
77 / 81 correct responses



Which Test is used to determine whether a machine can think?

Copy

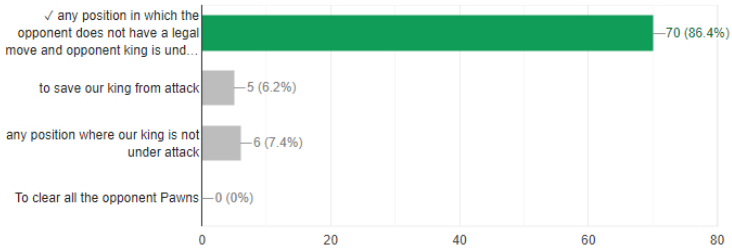
81 / 81 correct responses



In Playing Chess, Goal Position is _____.

Copy

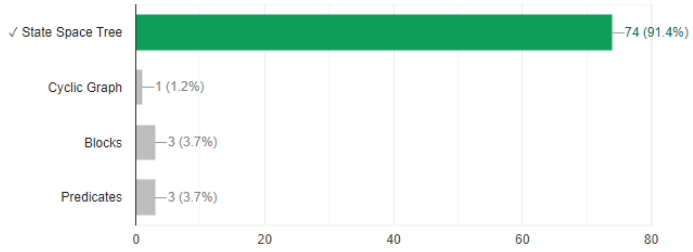
70 / 81 correct responses



Water-Jug Problem solution is represented using

Copy

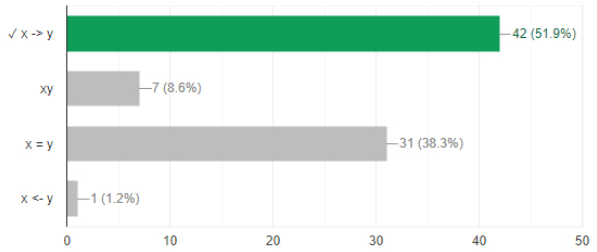
74 / 81 correct responses



Rules of the production system are of form :

Copy

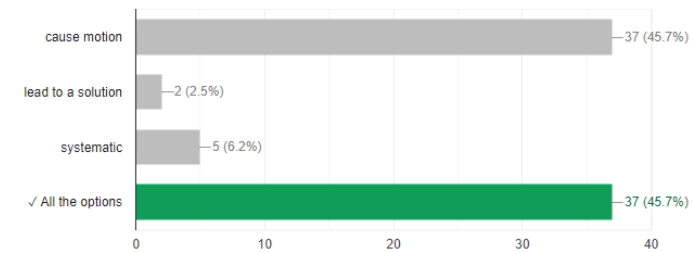
42 / 81 correct responses



A good control strategy must

Copy

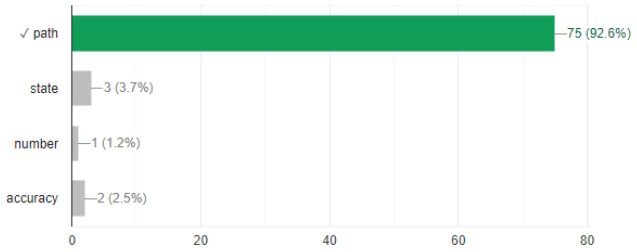
37 / 81 correct responses



The solution of the Travelling Salesman Problem is about finding a _____.

Copy

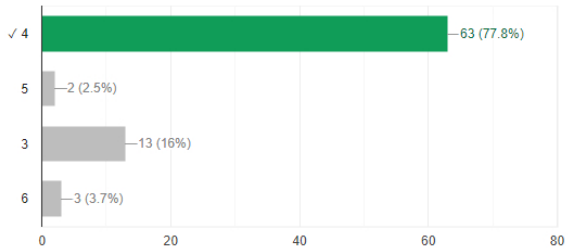
75 / 81 correct responses



A good system for the representation of knowledge in a particular domain should possess how many properties

Copy

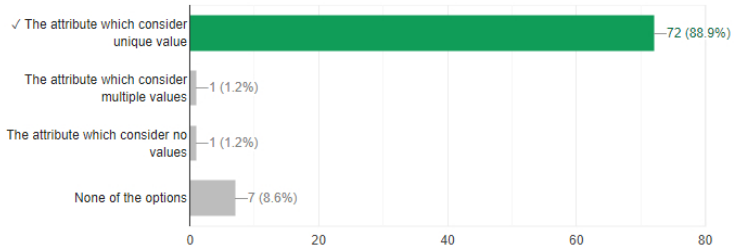
63 / 81 correct responses



Single value attributes means

Copy

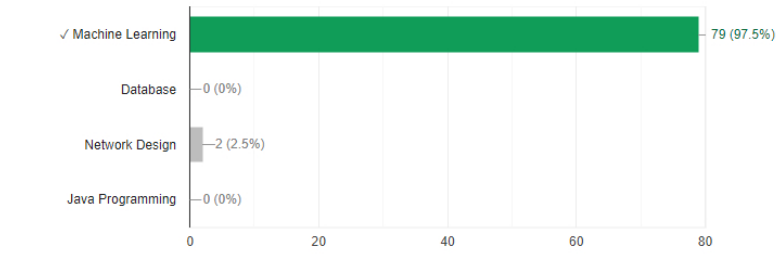
72 / 81 correct responses



Which one is considered as branch of Artificial Intelligence?

Copy

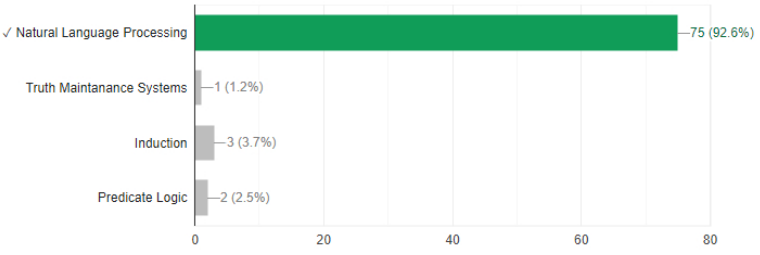
79 / 81 correct responses



English sentences grammar checking comes under

Copy

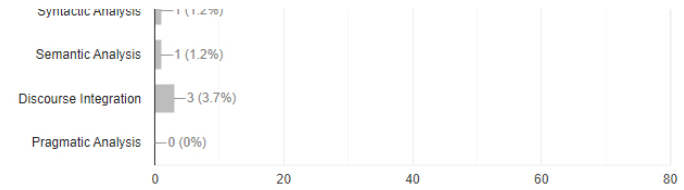
75 / 81 correct responses



Which is the First Stage in Natural Language Processing

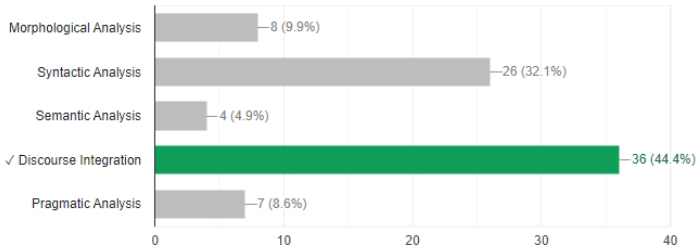
Copy

76 / 81 correct responses



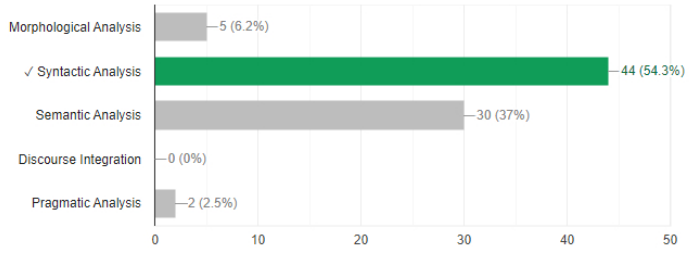
Which Stage of Natural Language Processing refer to preceding sentences? Copy

36 / 81 correct responses



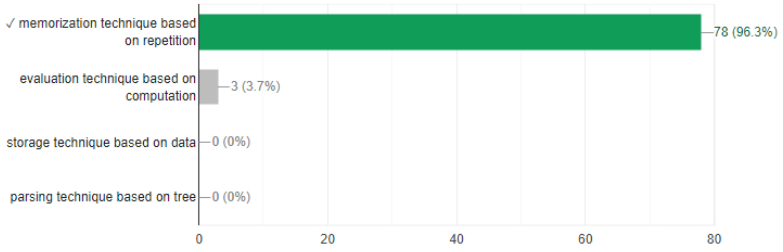
In _____ linear sequences of words are transformed into structures. Copy

44 / 81 correct responses



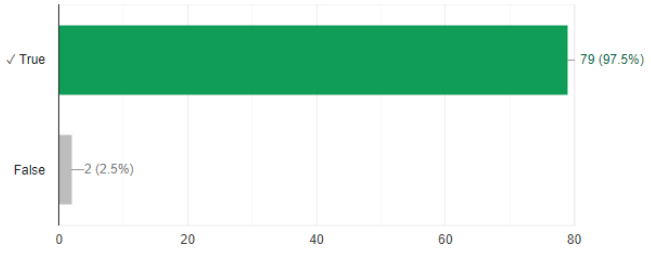
Rote learning is a _____. Copy

78 / 81 correct responses



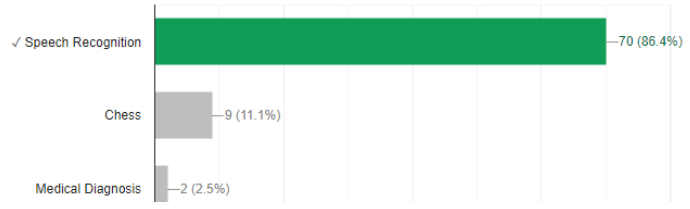
Intelligence requires knowledge Copy

79 / 81 correct responses



Which one is a Mundane Task? Copy

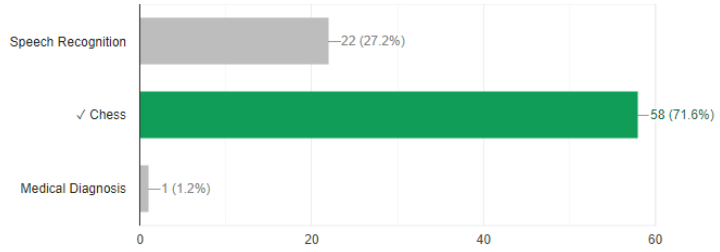
70 / 81 correct responses



Which one is a Formal Task?

Copy

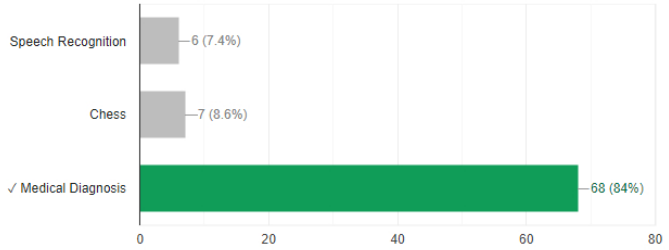
58 / 81 correct responses



Identify the Expert Task

Copy

68 / 81 correct responses



National Education Society ®

JAWAHARLAL NEHRU NEW COLLEGE OF ENGINEERING

Shivamogga - 577204

DEPARTMENT OF CIVIL ENGINEERING

PEDAGOGICAL INITIATIVE REPORT

18CV62 – APPLIED GEOTECHNICAL ENGINEERING

Quiz Assignment (Academic Year 2022-23)

6th Sem – Batch 2020-24

Topic – Pile Foundation

The students of 6 semester were given the quiz assignment on Module 5 (Pile Foundation). The quiz was conducted through google classroom on 29-06-2023.

The following are the questions covered in quiz.

Quiz Link: https://docs.google.com/forms/d/e/1FAIpQLScRJ-9TF_FQNbrNoR-rT2zQTZL326_m6jPIRm-eXRnRNFZk9Q/viewform?authuser=0

1. **Which of the following is not a characterstics of friction piles**

a. The majority amount of load is tranferred through skin friction between soil and pile

b. Transmit the load through weak soil to the hard stratum

c. The ultimate load carried by the pile is mainly due to the load transferref by skin friction.

d. These types of pile generally do not reach hard stratum.

Ans: 2

2. **Which type of pile is most suitable for light loads in soft cohesive soil**

a. Concrete Piles
b. Steel piles
c. Timber Piles
d. Both Concrete and steel piles

Ans: c

3. **A group of 16 piles of 600mm diameter is arranged in a square pattern with centre to centre spacing of 1.2m. the piles are 10m long and are embedded in soft clay with cohesion of 30kN/m^2. Bearing resistance may be neglected for the piles. Adhesion factor is 0.6. Determine the ultimate load capacity of pile group.**

a. 5428.7kN
b. 5040kN
c. 6040kN
d. 4428kN

Ans: b. 5040

4. **A group of 9 piles with 3 piles in a row were driven into a soft clay extending from ground level to a great depth. The diameter and length of the piles were 30cm and 10m respectively. The unconfined compressive strength of the clay is 70kPa. If the piles were placed at 90cm centre to centre, compute the allowable load on the pile group. Assume F =2.5**

a. 11740kN
b. 3168kN
c. 1267kN
d. 4340kN

Ans: c. 1267kN

5. **A 12m long, 30mm diameter pile is driven in uniform deposit of sand with angle of internal friction 40. The water table is at great depth. The average dry unit weight of sand is 18kN/m^3. Using N_q=137, calculate the safe load capacity of single pile with a f=2.5, K =1.5 and angle of wall friction is 30.**

a. 2091.7kN
b. 1057.81kN
c. 3149.5kN
d. 1259.8kN

Ans: d. 1259.8

6. **A precast concrete pile of size 50 X 50cm is to be driven into clay strata whose unconfined compressive strength is 220kN/m^2 and adhesion factor is 0.6. the length of the pile required to carry safe working load of 450kN with factor of safety 2.5 is**

a. 6.647m
b. 2. 10.23m
c. 2m
d. 3.3m

ANS: a. 6.647m

7. **Under Reamed piles are generally**

a. Driven piles
b. Bored piles
c. Precast piles
d. All of the above

Ans: b. Bored Piles

8. **In the pile foundation, which type of pile acts as columns and transmit the load through weak soil to a firm stratum at a greater depth.**

a. Footing piles
b. End bearing Piles
c. Friction Piles
d. Compaction Piles

Ans: b. End bearing Piles

9. **A square pile of section 30 X 30cm and length 10m penetrates a deposit of clay having c =50kN/m^2 and adhesion factor 0.8. The load carried by the pile shaft only is**

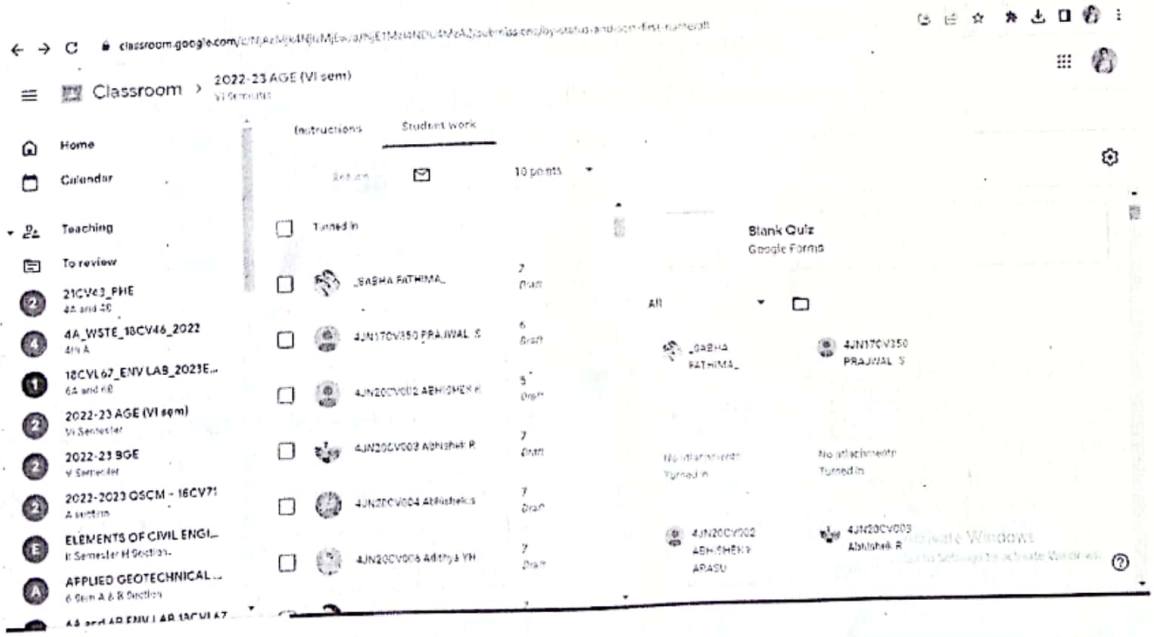
a. 1920kN
b. 750kN
c. 600kN
d. 480kN

Ans: d. 480kN

10. **Piles are structural members made of**

a. Steel
b. Concrete
c. Timber
d. All of the above

Ans: d. All of the above

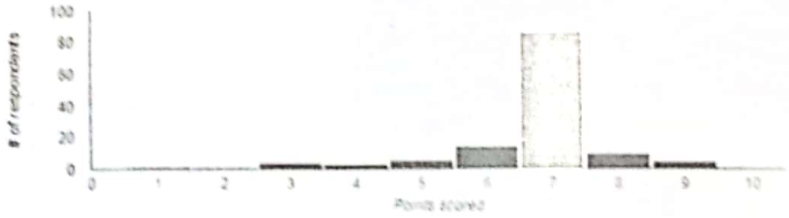


Email	Score / 10	Score released
bindiyakn@gmail.com	7	Jul 3 2:36 PM
echaran133@gmail.com	6	Jul 3 2:36 PM
akashkarthalkarthal@gmail.com	7	Jul 3 2:36 PM
sujayjay66@gmail.com	3	Jul 3 2:36 PM
chiranjanajnce@gmail.com	5	Jul 3 2:36 PM
amruthakademane@gmail.com	5	Jul 3 2:36 PM
mlakamveeby@gmail.com	6	Jul 3 2:36 PM
shubhanrubbu17@gmail.com	7	Jul 3 2:36 PM

Insights

Average	Median	Range
6.71 / 10 points	7 / 10 points	1 - 9 points

Total points distribution



BINDIYA K
Assistant Professor
Course co-ordinator

HOD

Professor & Head
Department of Civil Engineering
J.N.N. College of Engineering,
Shivamogga-577 204.

National Education Society ®
JAWAHARLAL NEHRU NATIONAL COLLEGE OF ENGINEERING
Shivamogga - 577204
DEPARTMENT OF CIVIL ENGINEERING

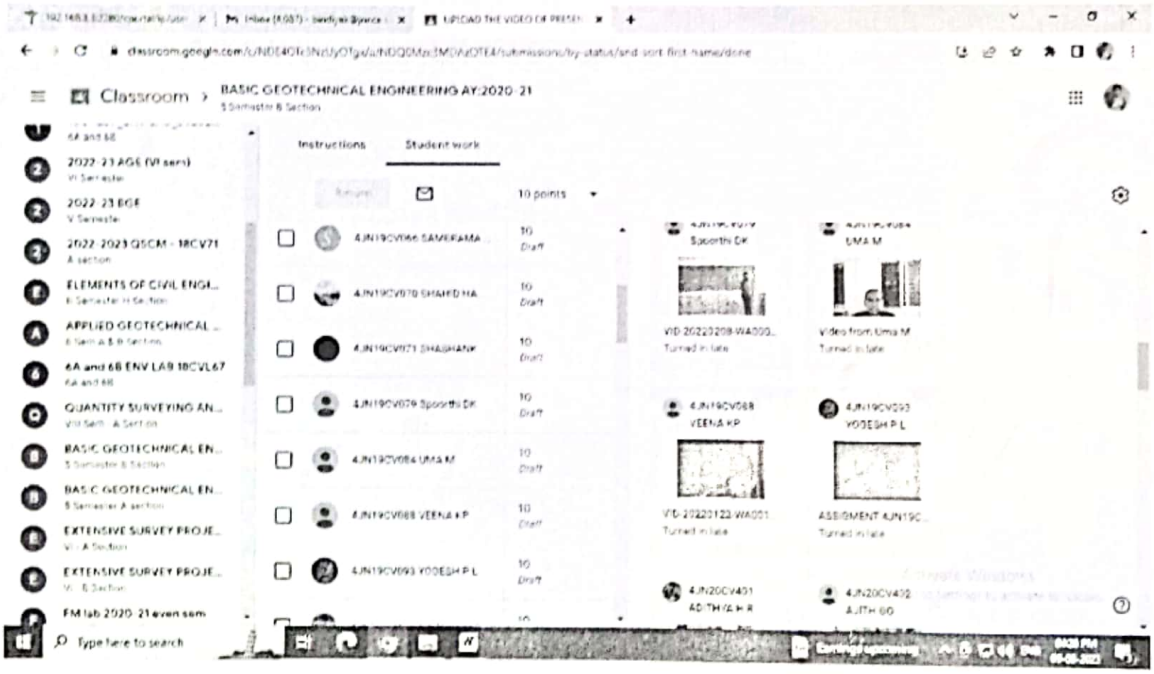
PEDAGOGICAL INITIATIVE REPORT

18CV54 – Basic Geotechnical Engineering
Video Assignment
5th Sem – 2021-22

Topic – Soil Structure and Clay Mineralogy

The students of 5 semesters were given the seminar topic of Soil Structure and Clay Mineralogy. They presented using power point presentation with chalk and board. They uploaded the videos in the Google classroom. They completed this activity before 10/01/2022.

Link:
https://classroom.google.com/c/NDE4OTc3NzUyOTgx/a/NDQ0Mzc3MDAzOTE4/details.



Bindiya K
Assistant Professor
Course Coordinator

Professor & Head
Department of Civil Engineering
J.N.N. College of Engineering,
Shivamogga-577 204.

Quizizz

Quiz Name	Date	Hosted by
corrosion and electrode system	Tue Jul 11 2023 3:51 AM	Chethan Chemistry

Average Accuracy	Questions per Attempt	Number of Players
81%	20	51

This report displays results derived from the students' best attempts.

Players

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
1	Kruthi B	21 secs	20	100%	20 / 20
2	Chethan JM	11 secs	20	100%	20 / 20
3	B.L.Maithreye B.L.MAITHREYE	10 secs	20	100%	20 / 20
4	INCHARA S	16 secs	20	100%	20 / 20
5	Jagath M D	27 secs	20	100%	20 / 20
6	Chethana T	10 secs	20	100%	20 / 20
7	Gaganpatil G V	22 secs	20	100%	20 / 20
8	Ananya.H Raj	16 secs	20	100%	20 / 20
9	Hima M	7 secs	20	100%	20 / 20
10	Abhijna S	19 secs	20	100%	20 / 20
11	Prajwal P	25 secs	20	100%	20 / 20
12	G .N.Nisarga Nisarga	27 secs	20	100%	20 / 20
13	Gagan Deep TV	34 secs	20	100%	20 / 20
14	Inchara S	6 secs	20	100%	20 / 20
15	Darshan Nv	31 secs	20	100%	20 / 20
16	Bhavana M K	6 secs	20	100%	20 / 20
17	Likhitha B S	4 secs	20	100%	20 / 20
18	Sp Kumar	34 secs	20	100%	20 / 20
19	Likhitha Likhitha	13 secs	20	100%	20 / 20
20	DEEKSHA M	20 secs	20	100%	20 / 20
21	Harsha K	31 secs	20	100%	20 / 20
22	Impana T.S.	25 secs	20	100%	20 / 20
23	Bhagath Mohan	14 secs	20	100%	20 / 20
24	Ananya P H	21 secs	20	100%	20 / 20
25	Inchara Poovaiah A	27 secs	20	100%	20 / 20
26	Hitashree S.G	11 secs	20	100%	20 / 20
27	Anvitha H M	26 secs	20	100%	20 / 20
28	Akshata Akshu	36 secs	20	100%	20 / 20
29	JYOTHI T N	21 secs	20	100%	20 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
30	Krupa U B	14 secs	19	95%	19 / 20
31	Harshitha S S	30 secs	19	95%	19 / 20
32	Kuberan A	9 secs	19	95%	19 / 20
33	dhvaneeth pbanakar	14 secs	19	95%	19 / 20
34	Likitha S	12 secs	19	95%	19 / 20
35	Amrutha M Holla	22 secs	19	95%	19 / 20
36	K P Nidhi	15 secs	19	95%	19 / 20
37	Uthpala R	74 secs	19	95%	19 / 20
38	Dhanalakshmi. S	24 secs	18	90%	18 / 20
39	Bhavana B M	94 secs	15	75%	15 / 20
40	Harsha Sangur	252 secs	13	65%	13 / 20
41	Gunjan Vjain	132 secs	13	65%	13 / 20
42	Dashami H N	188 secs	12	60%	12 / 20
43	Samruddhi Navale	203 secs	12	60%	12 / 20
44	Abhishek M N	3 secs	5	25%	5 / 20
45	Rishabh Jayaram	4 secs	5	25%	5 / 20
46	Aishwarya S	79 secs	2	10%	2 / 20
47	Ashwini S Ashu	248 secs	1	5%	1 / 20
48	Jnaneshwari R	0 secs	0	0%	0 / 20
49	HARSHA SB	18 secs	0	0%	0 / 20
50	HARSHA Sangur	0 secs	0	0%	0 / 20
51	Varun Varun	0 secs	0	0%	0 / 20

Quizizz

Quiz Name	Date	Hosted by
Corrosion	Fri Jul 02 2021 7:00 PM	chethan.s.g Sg

Average Accuracy	Questions per Attempt	Number of Players
64%	20	33

This report displays results derived from the students' best attempts.

Players

Rank	Player Name	Avg. Time	Score	Accuracy	Correct
1	4JN20ET036 Gujjar	19 secs	15210	100%	20 / 20
2	Varun V	11 secs	14750	85%	17 / 20
3	A M Mallikarjuna	13 secs	14840	85%	17 / 20
4	subhash chandra	8 secs	15540	85%	17 / 20
5	Meghana P Meghana Prabhakar	21 secs	12940	85%	17 / 20
6	Sannidhi TN	13 secs	13810	80%	16 / 20
7	Jeevan Jnnce	25 secs	11460	80%	16 / 20
8	Saishree P	13 secs	13650	80%	16 / 20
9	Aditi Sg	13 secs	13190	80%	16 / 20
10	shamita kamat	43 secs	11590	75%	15 / 20
11	Sneha Shidenur	6 secs	13590	75%	15 / 20
12	Nischitha r	30 secs	10180	70%	14 / 20
13	Ananya ETC	22 secs	10320	70%	14 / 20
14	Pranav swaroop K. S swaroop	12 secs	12290	70%	14 / 20
15	Bhoomika banakar	18 secs	11200	70%	14 / 20
16	Mahek Shaikh	15 secs	11590	70%	14 / 20
17	mohammed afroz	26 secs	10090	70%	14 / 20
18	Priyanka j	9 secs	12430	70%	14 / 20
19	SHIFANA FATHIMA-ETC	14 secs	10670	65%	13 / 20
20	Srinivas S	12 secs	11290	65%	13 / 20
21	Vismitha Salanke	19 secs	10210	65%	13 / 20
22	Tejashwini.D -TCE	17 secs	11050	65%	13 / 20
23	Rohini HN	29 secs	8980	65%	13 / 20
24	Inchara.J. Inchara.J.	8 secs	10610	60%	12 / 20
25	Bindu GP	19 secs	9240	60%	12 / 20
26	4JN20ET021 j	25 secs	8960	60%	12 / 20
27	Ankitha V	35 secs	8430	60%	12 / 20
28	Sudarshan Ramana	20 secs	8710	55%	11 / 20
29	Swathi C	7 secs	9080	50%	10 / 20

Rank	Player Name	Avg. Time	Score	Accuracy	Correct
30	Afham Baig	10 secs	8780	50%	10 / 20
31	Priyanka N	16 secs	7640	50%	10 / 20
32	Dharithri H L	6 secs	8320	45%	9 / 20
33	Hruthik E&Tc	3 secs	3780	20%	4 / 20

Quizizz

Quiz Name	Date	Hosted by
Electrochemistry	Mon Jun 28 2021 9:15 AM	chethan.s.g Sg

Average Accuracy	Questions per Attempt	Number of Players
62%	20	58

This report displays results derived from the students' best attempts.

Players

Rank	Player Name	Avg. Time	Score	Accuracy	Correct
1	Varsha C U	5 secs	23160	100%	20 / 20
2	Vachana Belgod-CSE	11 secs	22280	100%	20 / 20
3	Vanyashree JS	3 secs	22600	100%	20 / 20
4	Prajna Prakash	42 secs	18600	100%	20 / 20
5	Surya S	29 secs	19580	100%	20 / 20
6	Shreya S Bharadwaj	10 secs	18760	100%	20 / 20
7	Yashaswini I K	8 secs	20900	95%	19 / 20
8	Mohmmed Sahil	44 secs	15880	95%	19 / 20
9	Sagar SR	46 secs	18410	95%	19 / 20
10	Dhanyatha Gowda	45 secs	15650	90%	18 / 20
11	Jason W George	21 secs	15390	90%	18 / 20
12	Swathi R	35 secs	15180	90%	18 / 20
13	Suhas D.B	3 secs	16900	90%	18 / 20
14	Amrutha S	8 secs	16990	90%	18 / 20
15	samarth sammu	22 secs	17960	90%	18 / 20
16	Samarth ng	14 secs	17210	90%	18 / 20
17	Sanjana K. J	68 secs	16320	90%	18 / 20
18	Vibha GM	68 secs	13220	90%	18 / 20
19	Tanmayee Sharvani	12 secs	17700	85%	17 / 20
20	Shivansh V	13 secs	15200	85%	17 / 20
21	4JN20CS125- Ahmed	51 secs	16290	85%	17 / 20
22	Sufiyan Khan	5 secs	17120	85%	17 / 20
23	Sakshi Reddy	13 secs	17010	85%	17 / 20
24	Vineet Kowti	11 secs	17700	85%	17 / 20
25	VIJENDRA S.N	35 secs	14790	85%	17 / 20
26	Rukmini s S	8 secs	17510	85%	17 / 20
27	vishwas ss	27 secs	15880	85%	17 / 20
28	Junaidh Fardeen	8 secs	16150	80%	16 / 20
29	Vinayashree Shet	53 secs	14100	80%	16 / 20

Rank	Player Name	Avg. Time	Score	Accuracy	Correct
30	srijan sanicum	27 secs	14270	80%	16 / 20
31	Krupa H T Krupa H T	99 secs	12170	80%	16 / 20
32	Pooja R Pooja R	46 secs	11770	75%	15 / 20
33	Yathindra M	35 secs	14050	75%	15 / 20
34	Shravya D gowda	20 secs	14290	70%	14 / 20
35	Varshitha H	66 secs	10630	70%	14 / 20
36	Manya S H	45 secs	10890	70%	14 / 20
37	Syed Ahmed	55 secs	10830	70%	14 / 20
38	Anushree E C	10 secs	12150	65%	13 / 20
39	SUCHITRA R-CSE	70 secs	11490	65%	13 / 20
40	Ruchitha S R	75 secs	9560	65%	13 / 20
41	Sanjana T.R	9 secs	12090	65%	13 / 20
42	VAISHNAVI N	98 secs	10670	65%	13 / 20
43	Bhagannagouda P	53 secs	8770	60%	12 / 20
44	Shree -CSE	5 secs	10700	60%	12 / 20
45	SIRI K S	74 secs	9490	60%	12 / 20
46	Shreya Ganapathi	22 secs	11290	60%	12 / 20
47	Kushal KN	11 secs	10630	60%	12 / 20
48	Vinayaka Kudva	108 secs	10640	60%	12 / 20
49	4JN20CS026 Deepthi	63 secs	8090	55%	11 / 20
50	Pratheek CR	33 secs	8070	55%	11 / 20
51	MANIKYA G	6 secs	10400	50%	10 / 20
52	Banushree Sarji	110 secs	7440	50%	10 / 20
53	khushi raj	40 secs	7510	45%	9 / 20
54	Manoj P	56 secs	5330	35%	7 / 20
55	SANDEEP JADHAV	22 secs	4280	30%	6 / 20
56	Deepthi B.A	0 secs		0%	0 / 0
57	4JN20CS096 Shraddha C S Atreya	0 secs		0%	0 / 0
58	Spandana GL	0 secs		0%	0 / 0

Quizizz

Quiz Name	Date	Hosted by
Electrochemistry	Sun Jul 10 2022 7:54 PM	Chethan Chemistry

Average Accuracy	Total Questions	Number of Players	Participant Attempts
77%	20	78	142

This report displays results derived from the students' all attempts.

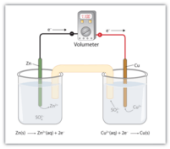
Questions

No.	Question	Time	Accuracy	Responses		
				Correct	*Incorrect*	*Unattempted*
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	28 secs	423%	120	6	16
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	32 secs	387%	110	14	18
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	19 secs	387%	110	13	19
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	23 secs	401%	114	8	20
5	An oxidizing agent will	19 secs	401%	114	11	17
6	As an element is oxidized, its oxidation number	31 secs	356%	101	25	16
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	17 secs	345%	98	28	16
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	22 secs	349%	99	24	19
9	Galvanic cells convert	17 secs	412%	117	9	16
10	When water is electrolyzed, gas collected at cathode, is	18 secs	415%	118	7	17
11	Conductivity always _____ with a decrease in concentration	17 secs	384%	109	12	21
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	13 secs	415%	118	4	20
13	Which are examples of reduction?	25 secs	370%	105	20	17
14	Which change does nitrogen undergo oxidation?	36 secs	387%	110	14	18
15	What reaction occurs at the anode?	21 secs	387%	110	14	18
16	Which direction do the electrons flow in wire X and which metal is oxidized?	27 secs	377%	107	16	19
17	Which statement best describes how	20 secs	377%	107	17	18

No.	Question	Time	Accuracy	Responses		
				Correct	*Incorrect*	*Unattempted*
	a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?					
18	When an electrochemical cell is operating, it is	34 secs	419%	119	4	19
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ E° = –0.44 volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ E° = –0.23 volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	27 secs	384%	109	15	18
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	33 secs	338%	96	29	17

Appendix - Images

1.

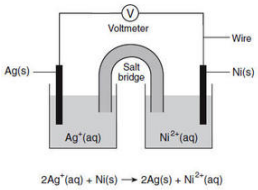


Given their standard reduction potentials, which of the species is going to be oxidized?

Cu^{2+}/Cu = 0.34V

Zn^{2+}/Zn = -0.76V

3.

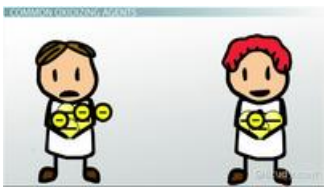


What reaction occurs at the anode?

Ag^+/Ag = 0.80V

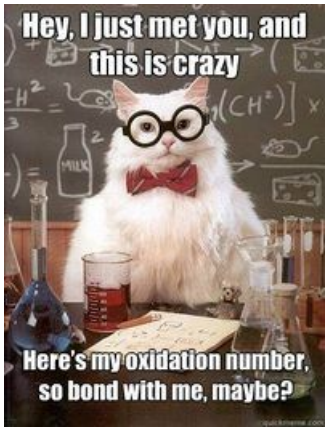
Ni^{2+}/Ni = -0.25V

5.



An oxidizing agent will

6.



As an element is oxidized, its oxidation number

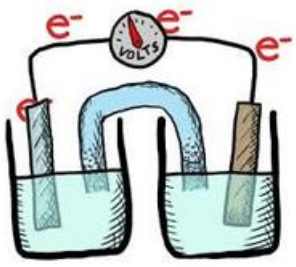
7.

In the following reaction $Sn^{+2} + 2Fe^{+3} \dashrightarrow Sn^{+4} + 2Fe^{+2}$, the reducing agent is...

8.

In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....

9.



Galvanic cells convert

10.

When water is electrolyzed, gas collected at cathode, is

13.

I. Fe^{3+} becomes Fe^{2+}

II. Cl^- becomes Cl_2

III. CrO_3 becomes Cr^{3+}

Which are examples of reduction?

14.

A. $NO_2 \rightarrow N_2O_4$

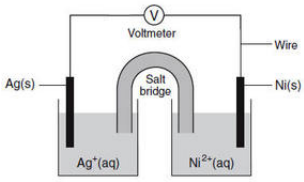
B. $NO_3^- \rightarrow NO_2$

C. $N_2O_5 \rightarrow NO_3^-$

D. $NH_3 \rightarrow N_2$

Which change does nitrogen undergo oxidation?

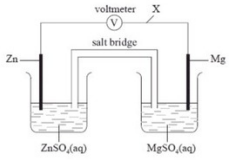
15.



$2Ag^+(aq) + Ni(s) \rightarrow 2Ag(s) + Ni^{2+}(aq)$

What reaction occurs at the anode?

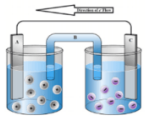
16.



	Electron flow	Oxidized
A.	Zn to Mg	Zn
B.	Mg to Zn	Zn
C.	Zn to Mg	Mg
D.	Mg to Zn	Mg

Which direction do the electrons flow in wire X and which metal is oxidized?

20.



Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?

Players

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
1	(N Nisarga)	61 secs	75	75%	15 / 20
2	(N Nisarga*)	6 secs	100	100%	20 / 20
3	(Vinay K M)	48 secs	85	85%	17 / 20
4	(Shashank HN)	38 secs	75	75%	15 / 20
5	(Shashank HN*)	10 secs	85	85%	17 / 20
6	(Pratiksha Shetty)	28 secs	70	70%	14 / 20
7	(Bhuvan)	46 secs	65	65%	13 / 20
8	(Bhuvan*)	8 secs	100	100%	20 / 20
9	(Bhuvan**)	6 secs	95	95%	19 / 20
10	(Mythri S P)	25 secs	95	95%	19 / 20
11	(Mythri S P*)	12 secs	15	15%	3 / 20
12	(Mythri S P**)	8 secs	65	65%	13 / 20
13	(Mythri S P***)	18 secs	0	0%	0 / 20
14	(Mythri S P****)	6 secs	100	100%	20 / 20
15	(Mythri S P*****)	4 secs	100	100%	20 / 20
16	(Nuthan S B)	63 secs	100	100%	20 / 20
17	(Saanvi BS)	51 secs	90	90%	18 / 20
18	(Saanvi BS*)	7 secs	65	65%	13 / 20
19	(Saanvi BS**)	8 secs	0	0%	0 / 20
20	(Saanvi BS***)	5 secs	100	100%	20 / 20
21	(PRAJWAL.KS)	32 secs	80	80%	16 / 20
22	(PRAJWAL.KS*)	7 secs	100	100%	20 / 20
23	(PRAJWAL.KS**)	5 secs	100	100%	20 / 20
24	(Prajwal DG)	0 secs	0	0%	0 / 20
25	(Prajwal DG*)	14 secs	40	40%	8 / 20
26	(Prajwal DG**)	6 secs	95	95%	19 / 20
27	(Prajwal DG***)	6 secs	100	100%	20 / 20
28	(Prajwal DG****)	3 secs	100	100%	20 / 20
29	(Anukeerthana MB)	29 secs	90	90%	18 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
30	(Anukeerthana MB*)	4 secs	55	55%	11 / 20
31	(Anukeerthana MB**)	6 secs	35	35%	7 / 20
32	(Anukeerthana MB***)	7 secs	100	100%	20 / 20
33	Deepa Shree (Deepashree.M*)	7 secs	100	100%	20 / 20
34	(Shubha H R)	12 secs	90	90%	18 / 20
35	(Raghu P R)	49 secs	95	95%	19 / 20
36	(Raghu P R*)	9 secs	95	95%	19 / 20
37	(Raghu P R**)	7 secs	100	100%	20 / 20
38	(Syeda Shafiya Anjum)	32 secs	80	80%	16 / 20
39	(Syeda Shafiya Anjum*)	4 secs	90	90%	18 / 20
40	(Syeda Shafiya Anjum**)	3 secs	95	95%	19 / 20
41	Priyanka Kadati (Priyanka Kadati)	6 secs	100	100%	20 / 20
42	(Chandana D R)	46 secs	80	80%	16 / 20
43	(Chandana D R*)	6 secs	90	90%	18 / 20
44	(Nanditha N Raj)	36 secs	70	70%	14 / 20
45	(Nanditha N Raj*)	7 secs	100	100%	20 / 20
46	(Revanth MA 4jn21cs129)	41 secs	100	100%	20 / 20
47	(Ananya.R)	67 secs	5	5%	1 / 20
48	(Ananya.R*)	0 secs	0	0%	0 / 20
49	(Ananya.R**)	21 secs	5	5%	1 / 20
50	(Sharanya Y S)	68 secs	100	100%	20 / 20
51	(Sharanya Y S*)	0 secs	0	0%	0 / 20
52	Pratheek T.G (Pratheek T.G*)	11 secs	80	80%	16 / 20
53	Pratheek T.G (Pratheek T.G**)	5 secs	100	100%	20 / 20
54	(Nisarga N)	74 secs	95	95%	19 / 20
55	(Nisarga N*)	7 secs	100	100%	20 / 20
56	(Shridhar BG 4JN21CS157)	45 secs	95	95%	19 / 20
57	(Bhoomika p)	53 secs	60	60%	12 / 20
58	(Bhoomika p*)	8 secs	80	80%	16 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
59	(Bhoomika p**)	5 secs	0	0%	0 / 20
60	(Rehan khan)	31 secs	100	100%	20 / 20
61	(SANGAM S S)	13 secs	45	45%	9 / 20
62	(SANGAM S S*)	7 secs	40	40%	8 / 20
63	(Rashmi K S)	65 secs	80	80%	16 / 20
64	(Rashmi K S*)	6 secs	95	95%	19 / 20
65	(Rashmi K S**)	4 secs	100	100%	20 / 20
66	(Bhavana v)	41 secs	55	55%	11 / 20
67	(Bhavana v*)	13 secs	100	100%	20 / 20
68	(Bhavana v**)	4 secs	100	100%	20 / 20
69	(Sanjay PS)	39 secs	95	95%	19 / 20
70	(Prateeksha A)	48 secs	100	100%	20 / 20
71	(wawa)	2 secs	15	15%	3 / 20
72	(wawa*)	2 secs	40	40%	8 / 20
73	(Nandan H.K.)	38 secs	70	70%	14 / 20
74	SAKETH N SHET (Saketh N Shet*)	10 secs	95	95%	19 / 20
75	(Poorvi T.C)	41 secs	75	75%	15 / 20
76	(Nayana HG)	66 secs	90	90%	18 / 20
77	(Nayana HG*)	0 secs	0	0%	0 / 20
78	(VN SUKUMAR)	45 secs	95	95%	19 / 20
79	(VN SUKUMAR*)	5 secs	100	100%	20 / 20
80	(Pallavi.g.v)	47 secs	90	90%	18 / 20
81	(Pallavi.g.v*)	8 secs	100	100%	20 / 20
82	(Pallavi.g.v**)	8 secs	100	100%	20 / 20
83	(Shrinidhi SR)	9 secs	100	100%	20 / 20
84	(Pramod J)	59 secs	80	80%	16 / 20
85	(Pramod J*)	7 secs	100	100%	20 / 20
86	(Shreya k .u)	33 secs	75	75%	15 / 20
87	(Shreya k .u*)	14 secs	95	95%	19 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
88	(Jogi)	0 secs	0	0%	0 / 20
89	(Saketh N Shet)	35 secs	95	95%	19 / 20
90	(PAREEKSHITH M)	9 secs	85	85%	17 / 20
91	(PAREEKSHITH M*)	6 secs	80	80%	16 / 20
92	(PAREEKSHITH M**)	4 secs	85	85%	17 / 20
93	(PAREEKSHITH M***)	4 secs	100	100%	20 / 20
94	(Patel M J)	36 secs	60	60%	12 / 20
95	(Patel M J*)	0 secs	0	0%	0 / 20
96	(Sahana k)	54 secs	100	100%	20 / 20
97	(Sahana k*)	9 secs	100	100%	20 / 20
98	(Nikhil BN)	33 secs	100	100%	20 / 20
99	(Sumanth p s)	29 secs	50	50%	10 / 20
100	(Sumanth p s*)	11 secs	100	100%	20 / 20
101	(Om Singh)	48 secs	100	100%	20 / 20
102	(Prathima H K)	59 secs	75	75%	15 / 20
103	(Shreya g)	15 secs	80	80%	16 / 20
104	(Mohammed Waseem)	7 secs	25	25%	5 / 20
105	(Mohammed Waseem*)	6 secs	65	65%	13 / 20
106	(S Nischal 4jn21cs131)	47 secs	80	80%	16 / 20
107	(Natasha. H. N.)	33 secs	85	85%	17 / 20
108	(Natasha. H. N.*)	4 secs	90	90%	18 / 20
109	(Natasha. H. N.**)	3 secs	100	100%	20 / 20
110	(Rajath s sajre)	0 secs	0	0%	0 / 20
111	(Swathi P Patel)	49 secs	80	80%	16 / 20
112	(Shreyas Karnik)	43 secs	85	85%	17 / 20
113	(Shreyas Karnik*)	14 secs	100	100%	20 / 20
114	(Punarvi B S)	23 secs	70	70%	14 / 20
115	(Shashwath K Rao)	22 secs	95	95%	19 / 20
116	(Shashwath K Rao*)	7 secs	100	100%	20 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
117	(Shashwath K Rao**)	6 secs	100	100%	20 / 20
118	(shami)	3 secs	100	100%	20 / 20
119	(Ñavaneeth Y)	58 secs	100	100%	20 / 20
120	(Ñavaneeth Y*)	12 secs	95	95%	19 / 20
121	(Rajath s sakre)	8 secs	85	85%	17 / 20
122	(Ruchitha K P)	48 secs	100	100%	20 / 20
123	(Ruchitha K P*)	0 secs	0	0%	0 / 20
124	(Nithin B M)	41 secs	95	95%	19 / 20
125	B. Nandan H K . Cse (Nandan H.K.*)	0 secs	0	0%	0 / 20
126	(Deepashree.M)	70 secs	100	100%	20 / 20
127	(Sathwik S M)	38 secs	95	95%	19 / 20
128	(Sathwik S M*)	5 secs	95	95%	19 / 20
129	(Ravikant shri Biradar)	45 secs	80	80%	16 / 20
130	(Ravikant shri Biradar*)	9 secs	85	85%	17 / 20
131	(Ravikant shri Biradar**)	9 secs	100	100%	20 / 20
132	(Prathiba vm)	43 secs	75	75%	15 / 20
133	(Pratheek T.G)	40 secs	100	100%	20 / 20
134	(Vaishnavi HK)	31 secs	75	75%	15 / 20
135	(Ananya R)	4 secs	100	100%	20 / 20
136	(Pranathi T)	63 secs	90	90%	18 / 20
137	(Pranathi T*)	11 secs	95	95%	19 / 20
138	(Priyanka Parashuram Kadat)	44 secs	100	100%	20 / 20
139	(Ankitha G S)	20 secs	35	35%	7 / 20
140	(Monisha y c)	31 secs	85	85%	17 / 20
141	Nithin Padthare (Nithin B M*)	5 secs	85	85%	17 / 20
142	(Shivani U)	56 secs	90	90%	18 / 20

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : N Nisarga (N Nisarga*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Bhuvan (Bhuvan*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Mythri S P (Mythri S P****)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nuthan S B (Nuthan S B)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Saanvi BS (Saanvi BS***)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : PRAJWAL.KS (PRAJWAL.KS**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Prajwal DG (Prajwal DG***)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Anukeerthana MB (Anukeerthana MB***)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Deepa Shree (Deepashree.M*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Raghu P R (Raghu P R**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Priyanka Kadati (Priyanka Kadati)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nanditha N Raj (Nanditha N Raj*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Revanth MA 4jn21cs129 (Revanth MA 4jn21cs129)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Sharanya Y S (Sharanya Y S)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pratheek T.G (Pratheek T.G**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nisarga N (Nisarga N*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Rehan khan (Rehan khan)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Rashmi K S (Rashmi K S**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Bhavana v (Bhavana v**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Prateeksha A (Prateeksha A)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : VN SUKUMAR (VN SUKUMAR*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pallavi.g.v (Pallavi.g.v**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shrinidhi SR (Shrinidhi SR)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pramod J (Pramod J*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : PAREEKSHITH M (PAREEKSHITH M***)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Sahana k (Sahana k)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nikhil BN (Nikhil BN)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Sumanth p s (Sumanth p s*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Om Singh (Om Singh)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Natasha. H. N. (Natasha. H. N.**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shreyas Karnik (Shreyas Karnik*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shashwath K Rao (Shashwath K Rao*)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : shami (shami)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ñavaneeth Y (Ñavaneeth Y)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ruchitha K P (Ruchitha K P)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Deepashree.M (Deepashree.M)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ravikant shri Biradar (Ravikant shri Biradar**)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pratheek T.G (Pratheek T.G)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ananya R (Ananya R)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Priyanka Parashuram Kadat (Priyanka Parashuram Kadat)

Accuracy

100%

Total Questions	✓ Correct	× Incorrect
20	20	0

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Syeda Shafiya Anjum (Syeda Shafiya Anjum**)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shridhar BG 4JN21CS157 (Shridhar BG 4JN21CS157)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Sanjay PS (Sanjay PS)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : SAKETH N SHET (Saketh N Shet*)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shreya k .u (Shreya k .u*)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Saketh N Shet (Saketh N Shet)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nithin B M (Nithin B M)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Sathwik S M (Sathwik S M)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pranathi T (Pranathi T*)

Accuracy

95%

Total Questions	✓ Correct	× Incorrect
20	19	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shubha H R (Shubha H R)

Accuracy

90%

Total Questions	✓ Correct	× Incorrect
20	18	2

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Chandana D R (Chandana D R*)

Accuracy

90%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	18	1	1

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nayana HG (Nayana HG)

Accuracy

90%

Total Questions	✓ Correct	× Incorrect
20	18	2

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shivani U (Shivani U)

Accuracy

90%

Total Questions	✓ Correct	× Incorrect
20	18	2

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Vinay K M (Vinay K M)

Accuracy

85%

Total Questions	✓ Correct	× Incorrect
20	17	3

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shashank HN (Shashank HN*)

Accuracy

85%

Total Questions	✓ Correct	× Incorrect
20	17	3

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Rajath s sakre (Rajath s sakre)

Accuracy

85%

Total Questions	✓ Correct	× Incorrect
20	17	3

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Monisha y c (Monisha y c)

Accuracy

85%

Total Questions	✓ Correct	× Incorrect
20	17	3

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nithin Padthare (Nithin B M*)

Accuracy

85%

Total Questions	✓ Correct	× Incorrect
20	17	3

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Bhoomika p (Bhoomika p*)

Accuracy

80%

Total Questions	✓ Correct	× Incorrect
20	16	4

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Shreya g (Shreya g)

Accuracy

80%

Total Questions	✓ Correct	× Incorrect
20	16	4

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : S Nischal 4jn21cs131 (S Nischal 4jn21cs131)

Accuracy

80%

Total Questions	✓ Correct	× Incorrect
20	16	4

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Swathi P Patel (Swathi P Patel)

Accuracy

80%

Total Questions	✓ Correct	× Incorrect
20	16	4

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Poorvi T.C (Poorvi T.C)

Accuracy

75%

Total Questions	✓ Correct	× Incorrect
20	15	5

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Prathima H K (Prathima H K)

Accuracy

75%

Total Questions	✓ Correct	× Incorrect
20	15	5

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Prathiba vm (Prathiba vm)

Accuracy

75%

Total Questions	✓ Correct	× Incorrect
20	15	5

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Vaishnavi HK (Vaishnavi HK)

Accuracy

75%

Total Questions	✓ Correct	× Incorrect
20	15	5

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Pratiksha Shetty (Pratiksha Shetty)

Accuracy

70%

Total Questions	✓ Correct	× Incorrect
20	14	6

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Nandan H.K. (Nandan H.K.)

Accuracy

70%

Total Questions	✓ Correct	× Incorrect
20	14	6

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Punarvi B S (Punarvi B S)

Accuracy

70%

Total Questions	✓ Correct	× Incorrect
20	14	6

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Mohammed Waseem (Mohammed Waseem*)

Accuracy

65%

Total Questions	✓ Correct	× Incorrect
20	13	7

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Patel M J (Patel M J)

Accuracy

60%

Total Questions	✓ Correct	× Incorrect
20	12	8

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : SANGAM S S (SANGAM S S)

Accuracy

45%

Total Questions	✓ Correct	× Incorrect
20	9	11

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : wawa (wawa*)

Accuracy

40%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	8	1	11

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ankitha G S (Ankitha G S)

Accuracy

35%

Total Questions	✓ Correct	× Incorrect
20	7	13

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Ananya.R (Ananya.R)

Accuracy

5%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	1	3	16

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Jogi (Jogi)

Accuracy

0%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	0	0	20

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : Rajath s sajre (Rajath s sajre)

Accuracy

0%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	0	0	20

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz : Electrochemistry

Date : Sun Jul 10 2022 7:54 PM

Student : B. Nandan H K . Cse (Nandan H.K.*)

Accuracy

0%

Total Questions	✓ Correct	× Incorrect	⊘ Unattempted
20	0	0	20

No.	Question	Time	Points	Response
1	Given their standard reduction potentials, which of the species is going to be oxidized? $Cu^{2+}/Cu = 0.34V$ $Zn^{2+}/Zn = -0.76V$	5	5	✓ Zn
2	What occurs to the mass of copper electrode in the following reaction? Zn/Zn^{2+} // Cu^{2+}/Cu	3	5	✓ increases
3	What reaction occurs at the anode? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	13	0	✗ $Ni^{2+} + 2e^{-} \rightarrow Ni$
4	What would be the theoretical cell potential of the previous electrochemical cell? $Ag^{+}/Ag = 0.80V$ $Ni^{2+}/Ni = -0.25V$	18	5	✓ 1.05V
5	An oxidizing agent will	5	5	✓ be reduced
6	As an element is oxidized, its oxidation number	22	5	✓ increases as electrons are lost
7	In the following reaction $Sn^{+2} + 2Fe^{+3}$ --> $Sn^{+4} + 2Fe^{+2}$, the reducing agent is...	10	5	✓ Sn^{+2}
8	In the following reaction $Sn^{+2} + 2Fe^{+3} \rightarrow Sn^{+4} + 2Fe^{+2}$ the oxidizing agent is....	9	5	✓ Fe^{+3}
9	Galvanic cells convert	4	5	✓ chemical energy in to electrical energy
10	When water is electrolyzed, gas collected at cathode, is	49	0	✗ sulphur
11	Conductivity always _____ with a decrease in concentration	34	5	✓ decreases
12	What is oxidation number of Cr in $Cr_2O_7^{2-}$?	3	5	✓ +6
13	Which are examples of reduction?	6	0	✗ I and II
14	Which change does nitrogen undergo oxidation?	135	5	✓ D
15	What reaction occurs at the anode?	14	5	✓ $Ni \rightarrow Ni^{2+} + 2e^{-}$
16	Which direction do the electrons flow in wire X and which metal is oxidized?	11	5	✓ D

No.	Question	Time	Points	Response
17	Which statement best describes how a salt bridge maintains electrical neutrality in the half-cells of an electrochemical cell?	19	5	✓ It permits the migration of ions.
18	When an electrochemical cell is operating, it is	11	5	✓ approaching equilibrium
19	$Fe^{2+} + 2e^- \rightarrow Fe(s)$ $E° = –0.44$ volt $Ni^{2+} + 2e^- \rightarrow Ni(s)$ $E° = –0.23$ volt The standard reduction potentials for two half reactions are given above. The Nernst equation for a galvanic cell at 25°C in which Fe(s) reduces Ni^{2+} is the following. $E = E° – 0.03 \log [Fe^{2+}]/[Ni^{2+}]$ What is the equilibrium constant for the reaction below? $Fe(s) + Ni^{2+} \rightarrow Fe^{2+} + Ni(s)$	30	5	✓ 1.3×10^7
20	Which of the following statements applies to the change in mass of the electrodes involved in this electrochemical cell?	8	5	✓ Electrode A is the cathode and it gains mass since metal ions are being converted to metal atoms which often adhere to the electrode.

Quizizz

Quiz Name	Date	Hosted by
Electrochemistry	Sun Jul 10 2022 7:54 PM	Chethan Chemistry

Average Accuracy	Questions per Attempt	Number of Players
77%	20	78

This report displays results derived from the students' best attempts.

Players

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
1	N Nisarga	6 secs	100	100%	20 / 20
2	Bhuvan	8 secs	100	100%	20 / 20
3	Mythri S P	6 secs	100	100%	20 / 20
4	Nuthan S B	63 secs	100	100%	20 / 20
5	Saanvi BS	5 secs	100	100%	20 / 20
6	PRAJWAL.KS	5 secs	100	100%	20 / 20
7	Prajwal DG	6 secs	100	100%	20 / 20
8	Anukeerthana MB	7 secs	100	100%	20 / 20
9	Deepa Shree	7 secs	100	100%	20 / 20
10	Raghu P R	7 secs	100	100%	20 / 20
11	Priyanka Kadati	6 secs	100	100%	20 / 20
12	Nanditha N Raj	7 secs	100	100%	20 / 20
13	Revanth MA 4jn21cs129	41 secs	100	100%	20 / 20
14	Sharanya Y S	68 secs	100	100%	20 / 20
15	Pratheek T.G	5 secs	100	100%	20 / 20
16	Nisarga N	7 secs	100	100%	20 / 20
17	Rehan khan	31 secs	100	100%	20 / 20
18	Rashmi K S	4 secs	100	100%	20 / 20
19	Bhavana v	4 secs	100	100%	20 / 20
20	Prateeksha A	48 secs	100	100%	20 / 20
21	VN SUKUMAR	5 secs	100	100%	20 / 20
22	Pallavi.g.v	8 secs	100	100%	20 / 20
23	Shrinidhi SR	9 secs	100	100%	20 / 20
24	Pramod J	7 secs	100	100%	20 / 20
25	PAREEKSHITH M	4 secs	100	100%	20 / 20
26	Sahana k	54 secs	100	100%	20 / 20
27	Nikhil BN	33 secs	100	100%	20 / 20
28	Sumanth p s	11 secs	100	100%	20 / 20
29	Om Singh	48 secs	100	100%	20 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
30	Natasha. H. N.	3 secs	100	100%	20 / 20
31	Shreyas Karnik	14 secs	100	100%	20 / 20
32	Shashwath K Rao	7 secs	100	100%	20 / 20
33	shami	3 secs	100	100%	20 / 20
34	Ñavaneeth Y	58 secs	100	100%	20 / 20
35	Ruchitha K P	48 secs	100	100%	20 / 20
36	Deepashree.M	70 secs	100	100%	20 / 20
37	Ravikant shri Biradar	9 secs	100	100%	20 / 20
38	Pratheek T.G	40 secs	100	100%	20 / 20
39	Ananya R	4 secs	100	100%	20 / 20
40	Priyanka Parashuram Kadat	44 secs	100	100%	20 / 20
41	Syeda Shafiya Anjum	3 secs	95	95%	19 / 20
42	Shridhar BG 4JN21CS157	45 secs	95	95%	19 / 20
43	Sanjay PS	39 secs	95	95%	19 / 20
44	SAKETH N SHET	10 secs	95	95%	19 / 20
45	Shreya k .u	14 secs	95	95%	19 / 20
46	Saketh N Shet	35 secs	95	95%	19 / 20
47	Nithin B M	41 secs	95	95%	19 / 20
48	Sathwik S M	38 secs	95	95%	19 / 20
49	Pranathi T	11 secs	95	95%	19 / 20
50	Shubha H R	12 secs	90	90%	18 / 20
51	Chandana D R	6 secs	90	90%	18 / 20
52	Nayana HG	66 secs	90	90%	18 / 20
53	Shivani U	56 secs	90	90%	18 / 20
54	Vinay K M	48 secs	85	85%	17 / 20
55	Shashank HN	10 secs	85	85%	17 / 20
56	Rajath s sakre	8 secs	85	85%	17 / 20
57	Monisha y c	31 secs	85	85%	17 / 20
58	Nithin Padthare	5 secs	85	85%	17 / 20

Rank	Player Name	Avg. Time	Points	Accuracy	Correct
59	Bhoomika p	8 secs	80	80%	16 / 20
60	Shreya g	15 secs	80	80%	16 / 20
61	S Nischal 4jn21cs131	47 secs	80	80%	16 / 20
62	Swathi P Patel	49 secs	80	80%	16 / 20
63	Poorvi T.C	41 secs	75	75%	15 / 20
64	Prathima H K	59 secs	75	75%	15 / 20
65	Prathiba vm	43 secs	75	75%	15 / 20
66	Vaishnavi HK	31 secs	75	75%	15 / 20
67	Pratiksha Shetty	28 secs	70	70%	14 / 20
68	Nandan H.K.	38 secs	70	70%	14 / 20
69	Punarvi B S	23 secs	70	70%	14 / 20
70	Mohammed Waseem	6 secs	65	65%	13 / 20
71	Patel M J	36 secs	60	60%	12 / 20
72	SANGAM S S	13 secs	45	45%	9 / 20
73	wawa	2 secs	40	40%	8 / 20
74	Ankitha G S	20 secs	35	35%	7 / 20
75	Ananya.R	67 secs	5	5%	1 / 20
76	Jogi	0 secs	0	0%	0 / 20
77	Rajath s sajre	0 secs	0	0%	0 / 20
78	B. Nandan H K . Cse	0 secs	0	0%	0 / 20